

**Table K
Selected Polyatomic Ions**

Formula	Name	Formula	Name
ClO_4^-	perchlorate	ClO_2^-	chlorite
ClO_3^-	chlorate	ClO_2^+	dichlorine oxide
NH_4^+	ammonium	MnO_4^-	permanganate
$\text{C}_2\text{H}_3\text{O}_2^-$ CH_3COO^-	acetate	NO_2^-	nitrite
CN^-	cyanide	NO_3^-	nitrate
CO_3^{2-}	carbonate	O_2^{2-}	peroxide
HCO_3^-	bicarbonate hydrogencarbonate	OH^-	hydroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate	PO_4^{3-}	phosphate
F_3B^-	borofluoride	SCN^-	thiocyanate
F_6O^-	hexafluoroantimonate	SO_3^{2-}	sulfite
F_6O_2^-	hexafluoroantimonate	SO_3^{2-}	sulfite
F_6O_3^-	hexafluoroantimonate	HSO_4^-	hydrogensulfate
F_6O_4^-	hexafluoroantimonate	$\text{S}_2\text{O}_8^{2-}$	persulfate

**Table L
Solubility Guidelines for Aqueous Solutions**

Basic Most Forms Soluble Compounds	Exceptions	Basic Most Forms Soluble Compounds	Exceptions
Group 1 ions (Li^+ , Na^+ , etc.)		acetate (CH_3CO_2^-)	when combined with Cu^{2+} or Pb^{2+}
ammonium (NH_4^+)		chromate (CrO_4^{2-})	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , or ammonium (NH_4^+)
nitrate (NO_3^-)		phosphate (PO_4^{3-})	when combined with Group 1 ions or ammonium (NH_4^+)
hydrogencarbonate (HCO_3^-)		oxide (O^{2-})	when combined with Group 1 ions or ammonium (NH_4^+)
halides (F^- , Cl^-)		hydroxide (OH^-)	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , or ammonium (NH_4^+)
halides (Br^- , I^-)	when combined with Ag^+ , Pb^{2+} , or Hg_2^{2+}		
sulfates (SO_4^{2-})	when combined with Ag^+ , Ca^{2+} , Sr^{2+} , Ba^{2+} , or Pb^{2+}		

*Sulfates with Group 1 ions are soluble in H_2O

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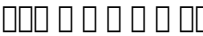
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