



Pork Chop Recipe Bone In

Megan Keno



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web oxidation reduction redox reactions oxidation $\text{Cu} \rightarrow \text{Cu}^{2+} + 2\text{e}^{-}$ reduction $\text{MnO}_4^{-} + 4\text{H}^{+} + 3\text{e}^{-} \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$ l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} : \text{MnO}_4^{-}$ is

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} \rightarrow \text{Au} + 2\text{e}^{-}$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 2\text{e}^{-}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} \rightarrow 2\text{Au} + 4\text{e}^{-}$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 2\text{e}^{-}$ 2oh 3 the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ and OH^- in this order balance the atoms in the equation apart from o and h

web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively

$\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$

choose 1 answer

$\text{C}_6\text{H}_8\text{O}_6$ and HNO_2

$\text{C}_6\text{H}_8\text{O}_6$ and H_2O

$\text{C}_6\text{H}_8\text{O}_6$ and NO

$\text{C}_6\text{H}_6\text{O}_6$ and HNO_2

$\text{C}_6\text{H}_6\text{O}_6$ and H_2O

$\text{C}_6\text{H}_6\text{O}_6$ and NO

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