

## Trends on the Periodic Table

### Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

**Lithium**  
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.

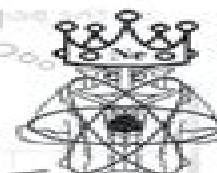


**Francium**  
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

**Fluorine**  
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (notorious electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



**Iodine**  
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

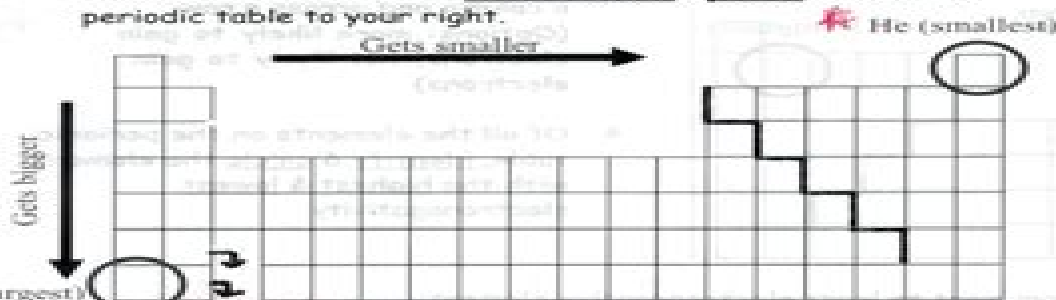


**Neon**  
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

### Part 2: Use your textbook to complete the following:

#### Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

*Effective nuclear charge,  $Z_{eff}$ , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.*

*not a good enough answer*  
**4AP Answer**

# Periodic Trends Pogil Answers Extension Questions

**S Baum**



### **Periodic Trends Pogil Answers Extension Questions:**

**Chemistry and Our Universe** , Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure Then witness what happens when three different alkali metals react with water Theory forecasts a pronounced difference in the result Is there      *Approximating Periodic and Non-periodic Trends in Time-series Data* Carlotta Ching Ting Fok,2002 The new model is then applied to Brown and Moskowitz s time series data to investigate the long term evolution to the four interpersonal behaviors and to the GDP data to examine the periodic and non periodic pattern for the GDP values of the 16 countries Finally the extent to which the model is accurate is tested using simulated data      Chemical Periodicity Robert Thomas Sanderson,2013-04-20

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web problems true or false the picture below is a resonance structure draw the lewis dot structure for  $\text{SO}_4^{2-}$  and all possible resonance structures which of the following

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web problem im5 1 draw lewis or kekule structures for the following molecules remembering to include formal charges if any and notice that some of these molecules are ions a

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web step 1 count the number of carbons and the number of hydrogens there are two carbons and six hydrogens a simple organic compound with no multiple bonds will always have

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web answer the following questions and check your answers below these problems are for practice only will not be graded be sure you know how to draw correct lewis dot

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web jul 20 2022 common bonding patterns in organic structures the methods reviewed above for drawing lewis structures and determining formal charges on atoms are an

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web 2 draw the lewis structures of the following polyatomic ions  $\text{SO}_2$   $\text{SO}_3^{2-}$   $\text{PO}_3^{3-}$   $\text{PO}_4^{3-}$   $\text{NO}_2^-$   $\text{NO}_3^-$   $\text{PF}_4^-$   $\text{PF}_6^-$   $\text{AlCl}_4^-$   $\text{AlCl}_3$

answer practise drawing the lewis

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web chapter 1 practice problems answers for these practice problems are on the next page a good approach is to answer all of the questions on a piece of paper and then check

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web two lewis structures labeled 1 and 2 in lewis structure 1 a central xenon with 2 lone pairs is line bonded to fluorines above and to the left and line bonded to chlorines below

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web lewis structures practice worksheet draw the lewis structures for each of the following molecules if you are not sure if your structure is correct do a formal

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web the skeletal structure of ethanethiol shows unbonded atoms a chain of two c atoms the first is surrounded by three h atoms above below and to the side the second is surrounded by two h atoms above and below and s atom attached to h atom to the side

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web species name lewis dot structure electronic arrangement molecular geometry bef 2 linear linear bcl 3 trigonal planar trigonal planar ccl 4 tetrahedral

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web 1 for all the reactants and products draw lewis structures 2 identify the nucleophile base and electrophile acid in the reaction 3 draw curved arrows to show the flow of

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web drawing lewis structures in short these are the steps you need to follow for drawing a lewis structure 1 write the correct skeletal structure for the molecule hydrogen