

CHAPTER 6 STUDY GUIDE

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, reads about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

octaves	atomic mass	atomic number	size
elements	properties	Henry Mosely	eight
protons	periodic law	Dmitri Mendeleev	<u>accepted</u>

The table below was developed by John Newlands and is based on a relationship called the law of (1) properties. According to this law, the properties of the elements repeated every (2) eight elements. Thus, for example, element two and element (3) nine have similar properties. The law of octaves did not work for all the known elements and was not generally (4) accepted.

1	2	3	4	5	6	7
H	Li	G	Be	C	N	O
8	9	10	11	12	13	15
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to (5) Mendeleev. In his table, the elements were arranged according to increasing (6) atomic mass. One important result of this table was that the existence and properties of undiscovered (7) elements could be predicted.

The element in the modern periodic table are arranged according to increasing (8) atomic number, as a result of the work of (9) Mosely. This arrangement is based on number of (10) protons in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) periodic law which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) properties.

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Table of Contents Periodic Law Answers To Remaining 16 Blocks

1. Understanding the eBook Periodic Law Answers To Remaining 16 Blocks
 - The Rise of Digital Reading Periodic Law Answers To Remaining 16 Blocks
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Law Answers To Remaining 16 Blocks
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Law Answers To Remaining 16 Blocks
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Law Answers To Remaining 16 Blocks
 - Personalized Recommendations

- Periodic Law Answers To Remaining 16 Blocks User Reviews and Ratings
- Periodic Law Answers To Remaining 16 Blocks and Bestseller Lists
- 5. Accessing Periodic Law Answers To Remaining 16 Blocks Free and Paid eBooks
 - Periodic Law Answers To Remaining 16 Blocks Public Domain eBooks
 - Periodic Law Answers To Remaining 16 Blocks eBook Subscription Services
 - Periodic Law Answers To Remaining 16 Blocks Budget-Friendly Options
- 6. Navigating Periodic Law Answers To Remaining 16 Blocks eBook Formats
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- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Law Answers To Remaining 16 Blocks
 - Highlighting and Note-Taking Periodic Law Answers To Remaining 16 Blocks
 - Interactive Elements Periodic Law Answers To Remaining 16 Blocks
- 8. Staying Engaged with Periodic Law Answers To Remaining 16 Blocks
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Law Answers To Remaining 16 Blocks
- 9. Balancing eBooks and Physical Books Periodic Law Answers To Remaining 16 Blocks
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Law Answers To Remaining 16 Blocks
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Law Answers To Remaining 16 Blocks
 - Setting Reading Goals Periodic Law Answers To Remaining 16 Blocks
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Law Answers To Remaining 16 Blocks
 - Fact-Checking eBook Content of Periodic Law Answers To Remaining 16 Blocks

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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