

Periodic Trends Worksheet

Name: _____ Grade: _____ ID: _____

Directions: Use your notes to answer the following questions.

- 1** Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
Ans: O, C, Al, K
- 2** Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
Ans: Ne, Al, S, O
- 3** Why does fluorine have a higher ionization energy than iodine?
Ans: Fluorine has a smaller atomic radius, so the proton can exert a greater pull on electron (Valence shell is close)
- 4** Why do elements in the same family generally have similar properties?
Ans: They have the same valence shell electron arrangement.
- 5** Indicate whether the following properties increase or decrease from left to right across the periodic table.
 - a. atomic radius (excluding noble gases) **Decrease**
 - b. first ionization energy **Increase**
 - c. electronegativity **Increase**
- 6** What trend in atomic radius occurs across the periodic table? What causes this trend?
Ans: Atomic radius increases down a group because energy levels (shells) are added.
- 7** What trend in ionization energy occurs across a period on the periodic table? What causes this trend?
Ans: Ionization energy increases across a period because as elements become less metallic, it requires more energy to remove an electron.
- 8** Circle the atom in each pair that has the largest radius.
 - a. Al or B
 - b. Na or Al
 - c. S or O
 - d. O or F
 - e. Br or Cl
 - f. Mg or Ca

Periodic Table Trends Answer Key

Albert A Gayle



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