Multiple Choice: Identify the letter of the choice that best completes the statement or answers the question.

- 1. What scientific (natural) law serves as the basis for balancing chemical equations by requiring that there be no observable change in the quantity of matter during a chemical reaction?
  - a. Law of Conservation of Energy
  - b. Law of Multiple Proportions
  - c. Law of Conservation of Matter
  - d. Law of Definite Proportions
  - e. Law of Constant Composition
- 2. Balance the following equation with the smallest whole number coefficients. What is the coefficient for O<sub>2</sub> in the balanced equation?

$$C_4H_{20} + O_2 \rightarrow CO_2 + H_2O$$

- a. 9
- b. 5
- c. 15
- d. 6
- e. 13
- 3. What is the coefficient for HBr when the following equation is balanced with the smallest whole number coefficients?

$$Br_2 + H_2O \rightarrow HBr + HBrO_3$$

- a. 5
- b. 7
- c. 8d. 3
- e. 6
- 4. Balance the following equation with the smallest whole number coefficients. What is the coefficient for NH<sub>3</sub> in the balanced equation?

$$Fe(NO_3)_3 + NH_3 + H_2O \rightarrow Fe(OH)_3 + NH_4NO_3$$

- a. 1b. 3
- b. 3 c. 2
- d. 6
- e. 4
- Elemental phosphorus is produced from calcium phosphate in the following reaction. What is the coefficient
  for C when this equation is balanced with the smallest whole number coefficients?

  Ca<sub>2</sub>(PO<sub>4</sub>)<sub>2</sub> + SiO<sub>2</sub> + C → P<sub>4</sub> + CO + CaSiO<sub>5</sub>
  - n. 10
  - b. 3
  - c. 1
  - d 6
  - e. 4

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