

(d) In an experiment, a student makes a solution of strontium chloride,  $\text{SrCl}_2$ , by adding excess dilute hydrochloric acid to strontium carbonate.

(i) Describe what the student would observe and write the equation for the reaction.

observations .....

.....

equation ..... [2]

(ii) Draw a 'dot-and-cross' diagram to show the bonding of strontium chloride. Show **outer** electrons only.

[2]

(e) In another experiment, a student attempts to make a solution of strontium chloride by adding chlorine water to aqueous strontium bromide.

(i) Describe what the student would observe.

..... [1]

(ii) Write the ionic equation for the reaction which takes place.

..... [1]

(iii) Chlorine is more reactive than bromine. Explain why.

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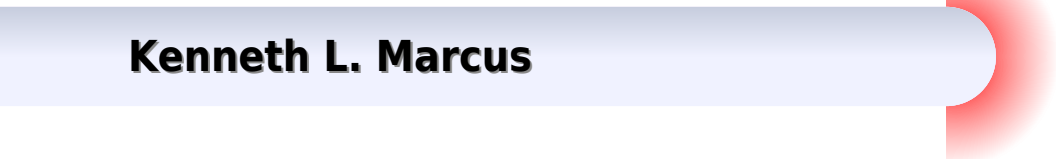
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