

Question	Answer	Mark	Comments
1	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$ (3 marks)	100 100	
2 (a)	Electron has longest wavelength (1000 nm)	100	100 mark bonus for diagram of diffraction
2 (b)	1000 nm	100	
2 (c)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100 100 100 100	
2 (d)	Electron has longest wavelength (1000 nm)	100	
2 (e)	Electron has longest wavelength (1000 nm)	100	
2 (f)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (g)	Electron has longest wavelength (1000 nm)	100	
2 (h)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (i)	Electron has longest wavelength (1000 nm)	100	
2 (j)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (k)	Electron has longest wavelength (1000 nm)	100	
2 (l)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (m)	Electron has longest wavelength (1000 nm)	100	
2 (n)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (o)	Electron has longest wavelength (1000 nm)	100	
2 (p)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (q)	Electron has longest wavelength (1000 nm)	100	
2 (r)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (s)	Electron has longest wavelength (1000 nm)	100	
2 (t)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (u)	Electron has longest wavelength (1000 nm)	100	
2 (v)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (w)	Electron has longest wavelength (1000 nm)	100	
2 (x)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	
2 (y)	Electron has longest wavelength (1000 nm)	100	
2 (z)	$E = hf$ $h = 6.63 \times 10^{-34} \text{ J s}$ $f = 5.0 \times 10^{14} \text{ Hz}$ $E = 3.315 \times 10^{-19} \text{ J}$	100 100	

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