

GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on a subshell
- readily react to lose 2 electrons
 - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^-$

Trends down the group:

- 1^{st} ionisation energy decreases
 - Distance between outermost electron and nucleus increases
 - More shielding
 - Counteracts effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
 - Much lower than all transition metals

Physical properties:

- Good conductors of heat and electricity
 - Sea of delocalised electrons
- Form white compounds
 - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are ionic or almost ionic
 - Be compounds show some covalent character because small highly charged cation polarises anion shell of anion
- Oxides / hydroxides are basic
 - React with acids (proton acceptors)
 - Except:
 - BeO
 - Be(OH)_2 which is amphoteric
- Density = $2.3\text{--}3.7\text{ g/cm}^3$
 - Average density of transition metals is about 7 g/cm^3
- Compared to group 2 metals:
 - Harder + denser
 - Higher melting point
 - Metallic bond is stronger (higher charge on cation $\therefore$ more electrons donated)

Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
 - Reducing agent - donates electrons (reducing other substances)
 - Oxidising agent - accepts electrons (oxidising itself)
- E.g.
 - $\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$
 - $\text{CaO} \rightarrow \text{CaO}^{2+} + 2\text{e}^-$
 - $\text{CO}_2 + 2\text{e}^- \rightarrow \text{CO}_3^{2-}$

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