

Chemistry - Unit 1 Chapters 1 & 2**Element Quiz:****Problem Set Due:****Vocab quiz:****Test Date:****Vocabulary:** Posted online – there will be a matching vocab quiz

- | | | | | | |
|------------------------|--------------------------|-------------------|-------------------|---------------|-------------|
| 1 matter | 2 temperature | 3 physical change | 4 chemical change | 5 precipitate | 6 chemistry |
| 7 compounds | 8 elements | 9 solution | 10 quantitative | 11 weight | 12 mass |
| 13 length | 14 SI | 15 accuracy | 16 precision | 17 density | 18 volume |
| 19 significant figures | 20 heterogeneous mixture | 21 endothermic | 22 exothermic | | |

Formulas/constants to be memorized:metric conversions given elements from handout
density(packet)

percent error (given)

Conversions that do NOT have to be memorized:

1 lb = 453.6 g 1 mile = 1.609 km 1 L = 1.0567 qts 1 oz = 28.35 g 4 qts = 1 gal
 1 m = 39.37 in = 3.2808 ft = 1.0936 yd (all equal! Use any 2 in a conversion) 1 fluid oz = 29.6 mL

OBJECTIVES:

- Be able to apply the concept of significant digits when recording lab data
- Memorize the metric system; be able to do both pure metric conversions and English-metric conversion. (You will be given English-metric conversions).
- Be able to calculate percent error from lab data.
- Use factor label (dimensional analysis) to solve math problems and conversions.
- Be able to do density problems.
- Be able to explain the difference between accuracy and precision and weight and mass.
- Be able to use the above vocabulary properly.
- Be able to classify properties and changes as physical or chemical.
- Understand and apply basic laboratory safety rules.
- Be able to visually identify and correctly use lab equipment and safety equipment.
- Know the fire escape route for the classroom and other safety procedures.
- Be able to collect, assimilate, and draw conclusions from scientific data.

Unit 1BQs We will work on this together in class! If you are absent you must still complete the BQs!! LABEL Q with numbers circle your final answers! Include units on your answers. 2. Show your work on all problems. Credit will not be given without work.

- Write the symbol or name for the following elements: K, Na, carbon, sulfur, nitrogen, lead and F
- In lab you find the mass of an object to be 5.0 g the correct mass is 5.2 g - calculate your percent error.
- Id the following as physical or chemical changes:
A. Wood burning in a bonfire B. inflating a tire with air C. ice cream cone melting D. sugar dissolving in water E. water absorbed by a paper towel F. your car using gasoline
- ID the following as heterogeneous mixtures, elements, solutions or compounds:
A. Kool-Aid B. ice C. pure gold coin D. vegetable soup E. diamond F. salt
- How many quarters in \$10?
- Convert: A. 55 meters to km B. 25 mm to cm C. 252 g to lb
- List two common sense lab safety rules.



- Record the length of the paper for A and B. which ruler is more accurate? Why? (the unit for both rulers is cm).
- What is the density of 20.0 g of a solution that has a volume of 38.2 ml?
- if you are traveling 25 ft/sec A. What is your speed in m/min ? B. How many km can you travel in half an hour?
- When a sample of a metal that has a mass of 15.8 g is placed into a graduated cylinder containing water, the volume increases from 16.0 ml to 19.5 ml.
A. What is the volume of the substance? B. What is the density of the metal?
C. Using your reference packet identify the metal.

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Disha Experts



Options For Youth English Packet Answers:

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Table of Contents Options For Youth English Packet Answers

1. Understanding the eBook Options For Youth English Packet Answers
 - The Rise of Digital Reading Options For Youth English Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Options For Youth English Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Options For Youth English Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Options For Youth English Packet Answers
 - Personalized Recommendations

- Options For Youth English Packet Answers User Reviews and Ratings
- Options For Youth English Packet Answers and Bestseller Lists
- 5. Accessing Options For Youth English Packet Answers Free and Paid eBooks
 - Options For Youth English Packet Answers Public Domain eBooks
 - Options For Youth English Packet Answers eBook Subscription Services
 - Options For Youth English Packet Answers Budget-Friendly Options
- 6. Navigating Options For Youth English Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Options For Youth English Packet Answers Compatibility with Devices
 - Options For Youth English Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Options For Youth English Packet Answers
 - Highlighting and Note-Taking Options For Youth English Packet Answers
 - Interactive Elements Options For Youth English Packet Answers
- 8. Staying Engaged with Options For Youth English Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Options For Youth English Packet Answers
- 9. Balancing eBooks and Physical Books Options For Youth English Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Options For Youth English Packet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Options For Youth English Packet Answers
 - Setting Reading Goals Options For Youth English Packet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Options For Youth English Packet Answers
 - Fact-Checking eBook Content of Options For Youth English Packet Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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