

11

CHEMICAL REACTIONS

Practice Problems

In your notebook, solve the following problems. Use the 3-step problem-solving approach you learned in Chapter 1.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS

1. Write the skeleton equation for the reaction between hydrogen and oxygen that produces water.
2. Write the skeleton equation for the reaction that produces iron(II) sulfide from iron and sulfur.
3. Write the skeleton equation representing the heating of magnesium carbonate to produce solid magnesium oxide and carbon dioxide gas.
4. Write a balanced equation for the production of HCl gas from its elements.
5. Write a sentence that completely describes the chemical reaction represented by this balanced equation.



6. Write the word equation for the following equation. Write a sentence fully describing the reaction. Is the equation correctly balanced? Explain.



7. Write a balanced equation representing the formation of aqueous sulfuric acid from water and sulfur trioxide gas.
8. Write a balanced equation from this word equation.
aqueous silver nitrate + copper metal \rightarrow silver metal + aqueous copper nitrate
9. Write a balanced equation for the following word equation.



SECTION 11.2 TYPES OF CHEMICAL REACTIONS

1. Write a balanced equation representing the reaction of magnesium with oxygen gas to produce magnesium oxide.
2. Write the balanced equation for the reaction that occurs between aluminum and fluorine.
3. Write the balanced equation for the production of oxygen gas and potassium chloride from the decomposition of potassium chlorate.
4. Write the balanced equation for the reaction between hydrochloric acid and calcium metal. The products are hydrogen gas and calcium chloride.
5. Write the balanced equation for the combustion of propane (C_3H_8) to produce carbon dioxide and water vapor.
6. Write the balanced equation for the reaction between iron(III) chloride and sodium hydroxide. The products are iron(III) hydroxide and sodium chloride.

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D Keegan

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Table of Contents Practice Problems Describing Chemical Reactions

1. Understanding the eBook Practice Problems Describing Chemical Reactions
 - The Rise of Digital Reading Practice Problems Describing Chemical Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Problems Describing Chemical Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Problems Describing Chemical Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Problems Describing Chemical Reactions

- Personalized Recommendations
 - Practice Problems Describing Chemical Reactions User Reviews and Ratings
 - Practice Problems Describing Chemical Reactions and Bestseller Lists
5. Accessing Practice Problems Describing Chemical Reactions Free and Paid eBooks
 - Practice Problems Describing Chemical Reactions Public Domain eBooks
 - Practice Problems Describing Chemical Reactions eBook Subscription Services
 - Practice Problems Describing Chemical Reactions Budget-Friendly Options
 6. Navigating Practice Problems Describing Chemical Reactions eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Problems Describing Chemical Reactions Compatibility with Devices
 - Practice Problems Describing Chemical Reactions Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Problems Describing Chemical Reactions
 - Highlighting and Note-Taking Practice Problems Describing Chemical Reactions
 - Interactive Elements Practice Problems Describing Chemical Reactions
 8. Staying Engaged with Practice Problems Describing Chemical Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Problems Describing Chemical Reactions
 9. Balancing eBooks and Physical Books Practice Problems Describing Chemical Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Problems Describing Chemical Reactions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Practice Problems Describing Chemical Reactions
 - Setting Reading Goals Practice Problems Describing Chemical Reactions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Practice Problems Describing Chemical Reactions

- Fact-Checking eBook Content of Practice Problems Describing Chemical Reactions
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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