

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES
(CHEMISTRY)**

TABLE 1 PHYSICAL CONSTANTS

NAME	SYMBOL	VALUE
Magnitude of charge on electron	e	$1,6 \times 10^{-19} \text{ C}$
Mass of an electron	m_e	$9,1 \times 10^{-31} \text{ kg}$
Standard pressure	p^θ	$1,01 \times 10^5 \text{ Pa}$
Molar gas volume at STP	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature	T^θ	273 K
Avogadro's constant	N_A	$6,02 \times 10^{23} \text{ mol}^{-1}$
Faraday's constant	F	$96\,500 \text{ C} \cdot \text{mol}^{-1}$

TABLE 2 CHEMISTRY FORMULAE

$n = \frac{m}{M}$	$n = \frac{N}{N_A}$	$n = \frac{V}{V_m}$
$c = \frac{n}{V}$ OR $c = \frac{m}{MV}$	$K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$ at 25°C (298 K)	
$q = It$ $q = nF$	$E_{\text{cell}}^\theta = E_{\text{cathode}}^\theta - E_{\text{anode}}^\theta$ $E_{\text{cell}}^\theta = E_{\text{oxidising agent}}^\theta - E_{\text{reducing agent}}^\theta$	

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TABLE OF CONTENTS Solved Paper 2019 Dec Solved Paper 2019 July Solved Paper 2018 Dec Solved Paper 2016 Sept Solved Paper 2016 Feb Practice Sets 1 15

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Table of Contents Physical Science P2 Grade1exampler

1. Understanding the eBook Physical Science P2 Grade1exampler
 - The Rise of Digital Reading Physical Science P2 Grade1exampler
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science P2 Grade1exampler
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science P2 Grade1exampler
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science P2 Grade1exampler
 - Personalized Recommendations

- Physical Science P2 Grade1exampler User Reviews and Ratings
- Physical Science P2 Grade1exampler and Bestseller Lists
- 5. Accessing Physical Science P2 Grade1exampler Free and Paid eBooks
 - Physical Science P2 Grade1exampler Public Domain eBooks
 - Physical Science P2 Grade1exampler eBook Subscription Services
 - Physical Science P2 Grade1exampler Budget-Friendly Options
- 6. Navigating Physical Science P2 Grade1exampler eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science P2 Grade1exampler Compatibility with Devices
 - Physical Science P2 Grade1exampler Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science P2 Grade1exampler
 - Highlighting and Note-Taking Physical Science P2 Grade1exampler
 - Interactive Elements Physical Science P2 Grade1exampler
- 8. Staying Engaged with Physical Science P2 Grade1exampler
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science P2 Grade1exampler
- 9. Balancing eBooks and Physical Books Physical Science P2 Grade1exampler
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science P2 Grade1exampler
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science P2 Grade1exampler
 - Setting Reading Goals Physical Science P2 Grade1exampler
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science P2 Grade1exampler
 - Fact-Checking eBook Content of Physical Science P2 Grade1exampler

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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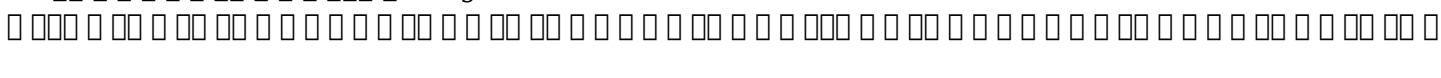
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