All quantities on the right are given to us except n, which can be computed from the given mass of Ar.

$$n = \frac{25 \text{ g}}{39.95 \text{ g mol}^{-1}} = 0.626 \text{ mol}$$

so
$$p = \frac{(0.626 \text{ mol}) \times (8.31 \times 10^{-2} \text{ dm}^3 \text{ bar } \text{K}^{-1} \text{ mol}^{-1}) \times (30 + 273 \text{ K})}{1.5 \text{ dm}^3} = 10.3 \text{ bar}$$

not 2.0 bar.

(b) The van der Waals equation [1.21a] is

$$p = \frac{RT}{V_m - b} - \frac{a}{V_m^2}$$

so
$$p = \frac{(8.31 \times 10^{-2} \,\mathrm{dm^3 \,bar \, K^{-1} \,mol^{-1}}) \times (30 + 273) \,\mathrm{K}}{(1.5 \,\mathrm{dm^3/0.626 \,mol}) - 3.20 \times 10^{-2} \,\mathrm{dm^3 \,mol^{-1}}}$$
$$- \frac{(1.337 \,\mathrm{dm^6 \,atm \,mol^{-2}}) \times (1.013 \,\mathrm{bar \,atm^{-1}})}{(1.5 \,\mathrm{dm^3/0.626 \,mol})^2} = \boxed{10.4 \,\mathrm{bar}}$$

E1.2(b) (a) Boyle's law [1.5] applies.

$$pV = constant$$
 so $p_tV_t = p_iV_i$

This equation can be solved for either initial or final pressure, hence

$$p_i = \frac{p_t V_f}{V_i} = \frac{(1.97 \text{ bar}) \times (2.14 \text{ dm}^3)}{(2.14 + 1.80) \text{ dm}^3} = [1.07 \text{ bar}]$$

(b) The original pressure in Torr is

$$p_i = (1.07 \text{ bar}) \times \left(\frac{1 \text{ atm}}{1.013 \text{ bar}}\right) \times \left(\frac{760 \text{ Torr}}{1 \text{ atm}}\right) = 803 \text{ Torr}$$

E1.3(b) The relationship between pressure and temperature at constant volume can be derived from the perfect gas law, pV = nRT[1.8]

so
$$p \approx T$$
 and $\frac{P_t}{T_t} = \frac{p_t}{T_t}$

The final pressure, then, ought to be

$$p_t = \frac{p_t T_t}{T_t} = \frac{(125 \text{ kPa}) \times (11 + 273) \text{ K}}{(23 + 273) \text{ K}} = \boxed{120 \text{ kPa}}$$

E1.4(b) According to the perfect gas law [1.8], one can compute the amount of gas from pressure, temperature, and volume.

$$pV = nRT$$

so
$$n = \frac{pV}{RT} = \frac{(1.00 \text{ atm}) \times (1.013 \times 10^5 \text{ Pa atm}^{-1}) \times (4.00 \times 10^5 \text{ m}^3)}{(8.3145 \text{ J K}^{-1} \text{ mol}^{-1}) \times (20 + 273) \text{ K}} = 1.66 \times 10^5 \text{ mol}$$

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Michael Rajnik

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 - The Rise of Digital Reading Physical Chemistry 9th Solution Manual
 - o Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Chemistry 9th Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Chemistry 9th Solution Manual
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Chemistry 9th Solution Manual
 - Personalized Recommendations
 - Physical Chemistry 9th Solution Manual User Reviews and Ratings
 - Physical Chemistry 9th Solution Manual and Bestseller Lists
- 5. Accessing Physical Chemistry 9th Solution Manual Free and Paid eBooks
 - Physical Chemistry 9th Solution Manual Public Domain eBooks
 - Physical Chemistry 9th Solution Manual eBook Subscription Services
 - Physical Chemistry 9th Solution Manual Budget-Friendly Options

- 6. Navigating Physical Chemistry 9th Solution Manual eBook Formats
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 - Physical Chemistry 9th Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Chemistry 9th Solution Manual
 - Highlighting and Note-Taking Physical Chemistry 9th Solution Manual
 - Interactive Elements Physical Chemistry 9th Solution Manual
- 8. Staying Engaged with Physical Chemistry 9th Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Chemistry 9th Solution Manual
- 9. Balancing eBooks and Physical Books Physical Chemistry 9th Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Chemistry 9th Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Chemistry 9th Solution Manual
 - Setting Reading Goals Physical Chemistry 9th Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Chemistry 9th Solution Manual
 - Fact-Checking eBook Content of Physical Chemistry 9th Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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