CHEM 1A

 Hydrogen has a red emission line at 656.3 nm, what is the energy and frequency of a photon of this light? Note: Planck's constant = 6.626 x 10⁻³⁴ J-s, speed of light = 2.998 x 10⁸ m/s

$$E_{photon} = hv = \frac{hc}{\lambda}$$

$$E_{photon} = \frac{(6.626 \times 10^{-34} \text{ J·s}) (2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left(\frac{10^{-9} \text{ m}}{1 \text{ nm}}\right)}$$

$$V = \frac{(2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left(\frac{10^{-9} \text{ m}}{1 \text{ nm}}\right)}$$

$$E_{photon} = \frac{3.02677861 \times 10^{-19} \text{ J}}{4 \text{ m/s}}$$

$$V = \frac{4.568032912 \times 10^{14} \text{ s}^{-1}}{4 \text{ m/s}}$$

Answer (E): 3.027 x 10⁻¹⁹ J

Answer (v): 4.568 x 10¹⁴ s⁻¹ or Hz

2. What is the de Broglie wavelength for an electron moving at 2.42 x 10° m/s? Note: m_e = 9.109 x 10⁻³¹ kg

$$\lambda = \left[\frac{b}{mei} \right]$$

$$\lambda = \left[\frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s}}{(9.109 \times 10^{-31} \text{ ke}) (2.42 \times 10^6 \text{ m/s})} \right] = 3.00583657 \times 10^{-10} \text{ m}$$

Answer: 3.01 x 10⁻¹⁰ m — many orders of magnitude larger than the diameter of an electron

 Using the Bohr equation, calculate is the change in energy when an electron transitions from n = 3 to n= 2 in a hydrogen atom.

$$\Delta E = -2.18 \times 10^{-18} J \left(\frac{Z^2}{n_{\text{final}}^2} - \frac{Z^2}{n_{\text{binal}}^2} \right)$$

$$\Delta E = -2.18 \times 10^{-18} J \left(\frac{1^2}{2^2} - \frac{1^2}{3^2} \right)$$

$$\Delta E = -3.02777778 \times 10^{-19} J$$

Answer: - 3.03 x 10⁻¹⁹ J

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CL Gary

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- 1. Understanding the eBook Practice Problems Chapter 7 Quantum Theory
 - The Rise of Digital Reading Practice Problems Chapter 7 Quantum Theory
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Practice Problems Chapter 7 Quantum Theory
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Problems Chapter 7 Quantum Theory
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Practice Problems Chapter 7 Quantum Theory
 - Personalized Recommendations

- Practice Problems Chapter 7 Quantum Theory User Reviews and Ratings
- Practice Problems Chapter 7 Quantum Theory and Bestseller Lists
- 5. Accessing Practice Problems Chapter 7 Quantum Theory Free and Paid eBooks
 - Practice Problems Chapter 7 Quantum Theory Public Domain eBooks
 - Practice Problems Chapter 7 Quantum Theory eBook Subscription Services
 - Practice Problems Chapter 7 Quantum Theory Budget-Friendly Options
- 6. Navigating Practice Problems Chapter 7 Quantum Theory eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Practice Problems Chapter 7 Quantum Theory Compatibility with Devices
 - Practice Problems Chapter 7 Quantum Theory Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Problems Chapter 7 Quantum Theory
 - Highlighting and Note-Taking Practice Problems Chapter 7 Quantum Theory
 - Interactive Elements Practice Problems Chapter 7 Quantum Theory
- 8. Staying Engaged with Practice Problems Chapter 7 Quantum Theory
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Problems Chapter 7 Quantum Theory
- 9. Balancing eBooks and Physical Books Practice Problems Chapter 7 Quantum Theory
 - ∘ Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Practice Problems Chapter 7 Quantum Theory
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Problems Chapter 7 Quantum Theory
 - Setting Reading Goals Practice Problems Chapter 7 Quantum Theory
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Problems Chapter 7 Quantum Theory
 - Fact-Checking eBook Content of Practice Problems Chapter 7 Quantum Theory

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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