

CHEM 1411 – General Chemistry I
Practice Problems, Chapters 1–3

Chapter 1 – Chemistry: The Study of Change

- 1.** Element, compound, homogeneous mixture (solution), or heterogeneous mixture:
- | | | |
|---------------------------------------|-----------------------|---|
| a) orange juice
(freshly-squeezed) | b) brass | c) 0.9% saline (NaCl) solution |
| d) garden soil | e) room air | f) methane gas |
| g) sodium metal | h) N ₂ gas | i) Cu(NO ₃) ₂ crystals |
- 2.** Define (some of these terms are found in Chapters 2 and 3):
- | | |
|-------------------------|--------------------------------|
| a) element | b) compound |
| c) solution | d) physical change |
| e) chemical change | f) law of conservation of mass |
| g) intensive properties | h) extensive properties |
| i) precision | j) accuracy |
| k) isotopes | l) stoichiometry |
| m) the mole | n) Avogadro's number |
| o) limiting reactant | |
- 3.** Convert
- | |
|--|
| a) 5 feet 10 inches to meters |
| b) 55 miles per hour to cm per second |
| c) 5.0 quarts to liters |
| d) 1.00 ft ² to cm ² |
| e) 32.06 amu to g |
| f) 25 g of carbon to moles of carbon |
| g) 1.00 g of gold to number of atoms of gold |
| h) 25 °F to Kelvins |
- 4.** How many significant figures are in the following numbers?
- | | | |
|-------------|--|---------------|
| a) 3.1416 | b) 6.022 × 10 ²³ atoms/mole | c) 0.000104 g |
| d) 6.0 g/mL | e) 1.00794 amu | f) 5000 mi |

Practice General Chemistry First Exam Asc

Rachel S Tattersall



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Table of Contents Practice General Chemistry First Exam Asc

1. Understanding the eBook Practice General Chemistry First Exam Asc
 - The Rise of Digital Reading Practice General Chemistry First Exam Asc
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice General Chemistry First Exam Asc
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice General Chemistry First Exam Asc
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice General Chemistry First Exam Asc
 - Personalized Recommendations
 - Practice General Chemistry First Exam Asc User Reviews and Ratings
 - Practice General Chemistry First Exam Asc and Bestseller Lists
5. Accessing Practice General Chemistry First Exam Asc Free and Paid eBooks

- Practice General Chemistry First Exam Asc Public Domain eBooks
 - Practice General Chemistry First Exam Asc eBook Subscription Services
 - Practice General Chemistry First Exam Asc Budget-Friendly Options
6. Navigating Practice General Chemistry First Exam Asc eBook Formats
- ePub, PDF, MOBI, and More
 - Practice General Chemistry First Exam Asc Compatibility with Devices
 - Practice General Chemistry First Exam Asc Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Practice General Chemistry First Exam Asc
 - Highlighting and Note-Taking Practice General Chemistry First Exam Asc
 - Interactive Elements Practice General Chemistry First Exam Asc
8. Staying Engaged with Practice General Chemistry First Exam Asc
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice General Chemistry First Exam Asc
9. Balancing eBooks and Physical Books Practice General Chemistry First Exam Asc
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice General Chemistry First Exam Asc
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Practice General Chemistry First Exam Asc
- Setting Reading Goals Practice General Chemistry First Exam Asc
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Practice General Chemistry First Exam Asc
- Fact-Checking eBook Content of Practice General Chemistry First Exam Asc
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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