

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

Physical Science 2nd Semester Exam Review

DP Hallahan



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web read philippians 4 12 13 paul says that he had learned the secret of being filled and going hungry both of having abundance and suffering need 4 12 that secret is stated in verse 13 i can do all things in him who continually infuses

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web oct 30 2019 philippians 4 10 13 and contentment i had an interesting talk with a man at church recently he shared with me that the first church he attended after accepting christ was a health and wealth type church if something bad happened in your life there was an assumption that you had sinned and were being punished for it

philippians 4 10 13 niv thanks for their gifts i rejoiced bible - Sep 05 2022

web thanks for their gifts 10 i rejoiced greatly in the lord that at last you renewed your concern for me indeed you were concerned but you had no opportunity to show it 11 i am not saying this because i am in need for i have learned to be content whatever the circumstances 12 i know what it is to be in need and i know what it is to have

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web mar 24 2019 series philippians finding joy in the gospel topic philippians scripture philippians 4 10 13 first scripture reading genesis 50 1 26 second scripture reading 1 timothy 6 3 21 sermon text philippians 4 10 13 we could think in terms of three main categories of discontentment with god's providence in our lives