



NATIONAL  
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2  
(CHEMISTRY)  
EXEMPLAR

MARKS: 100  
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.  
Use permanent marker for writing and drawing.

# QUESTION 3

## PHYSICAL SCIENCES PAPER 2

## INTERMOLECULAR FORCES

**GRADE 11**

**NOVEMBER 2020**

Examiner's use: \_\_\_\_\_ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the  $\Delta H_{\text{vap}}$  and  $\Delta H_{\text{fus}}$  for 1-hexanol in kJ mol<sup>-1</sup> and in  $^{\circ}\text{C}$  respectively.

Compound	Molecular mass (M <sub>r</sub> )	Boiling point (T <sub>b</sub> )
A	102	162
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 In what phase was compound A and B at 0  $^{\circ}\text{C}$  and 100  $^{\circ}\text{C}$  of boiling point? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What kind of the compound A and B is 1-hexanol? Explain how you arrived at your answer. (2)
- 3.1.5 Which compound (A or B) will have the higher vapour pressure? (1)
- 3.1.6 Give a reason for your answer. (1)
- 3.1.7 Which compound (A or B) will have the higher surface tension? (1)
- 3.1.8 Give a reason for the answer to 3.1.7 relating to the space the table. (1)

Compound B (M<sub>r</sub> 102) will show in the table has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 162  $^{\circ}\text{C}$ .

- 3.1.8 Explain why the boiling point of compound B is relatively higher than compound A and B by referring to the T<sub>b</sub> and molecular of intermolecular forces present. (2)

3.2 Substances C is added to water whereas D cannot be dissolved in water. In separate test tubes the above is the diagram below.



- 3.2.1 In which test tube (C or D), will the solute dissolve? (1)
- 3.2.2 Explain the answer to QUESTION 3.2.1 about the solubility in the intermolecular forces involved. (2)

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# Physical Science Paper2 November 2014

**Satabdi Das**



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