

# Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

**K, Na, O, S, Cl**

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

**Na, K, S, Cl, O**

3. Which atom has the higher electronegativity: sodium or sulfur?

**Sulfur (sulfur is to the right of sodium and closer to fluorine)**

4. Which two elements in the same family generally have similar properties?

**There have been two cases here where a strong periodic trend is present.**

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (decreasing to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

**Period 7, francium (francium is the furthest left and furthest down)**

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

**Period 2, fluorine (fluorine is the furthest right and furthest up)**

8. Write the atoms in each group that have the largest atomic radius.

a. **Li**  
 b. **Na**  
 c. **K**  
 d. **Rb**  
 e. **Cs**  
 f. **Fr**

a. **He**  
 b. **Ne**  
 c. **Ar**  
 d. **Kr**  
 e. **Xe**  
 f. **Rn**

9. Circle the atom in each group that has the greatest electronegativity.

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### **Table of Contents Physical Science If8767 Pg 63 Key**

1. Understanding the eBook Physical Science If8767 Pg 63 Key
  - The Rise of Digital Reading Physical Science If8767 Pg 63 Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Pg 63 Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science If8767 Pg 63 Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Pg 63 Key
  - Personalized Recommendations
  - Physical Science If8767 Pg 63 Key User Reviews and Ratings
  - Physical Science If8767 Pg 63 Key and Bestseller Lists

5. Accessing Physical Science If8767 Pg 63 Key Free and Paid eBooks
  - Physical Science If8767 Pg 63 Key Public Domain eBooks
  - Physical Science If8767 Pg 63 Key eBook Subscription Services
  - Physical Science If8767 Pg 63 Key Budget-Friendly Options
6. Navigating Physical Science If8767 Pg 63 Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science If8767 Pg 63 Key Compatibility with Devices
  - Physical Science If8767 Pg 63 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science If8767 Pg 63 Key
  - Highlighting and Note-Taking Physical Science If8767 Pg 63 Key
  - Interactive Elements Physical Science If8767 Pg 63 Key
8. Staying Engaged with Physical Science If8767 Pg 63 Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science If8767 Pg 63 Key
9. Balancing eBooks and Physical Books Physical Science If8767 Pg 63 Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science If8767 Pg 63 Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physical Science If8767 Pg 63 Key
  - Setting Reading Goals Physical Science If8767 Pg 63 Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science If8767 Pg 63 Key
  - Fact-Checking eBook Content of Physical Science If8767 Pg 63 Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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