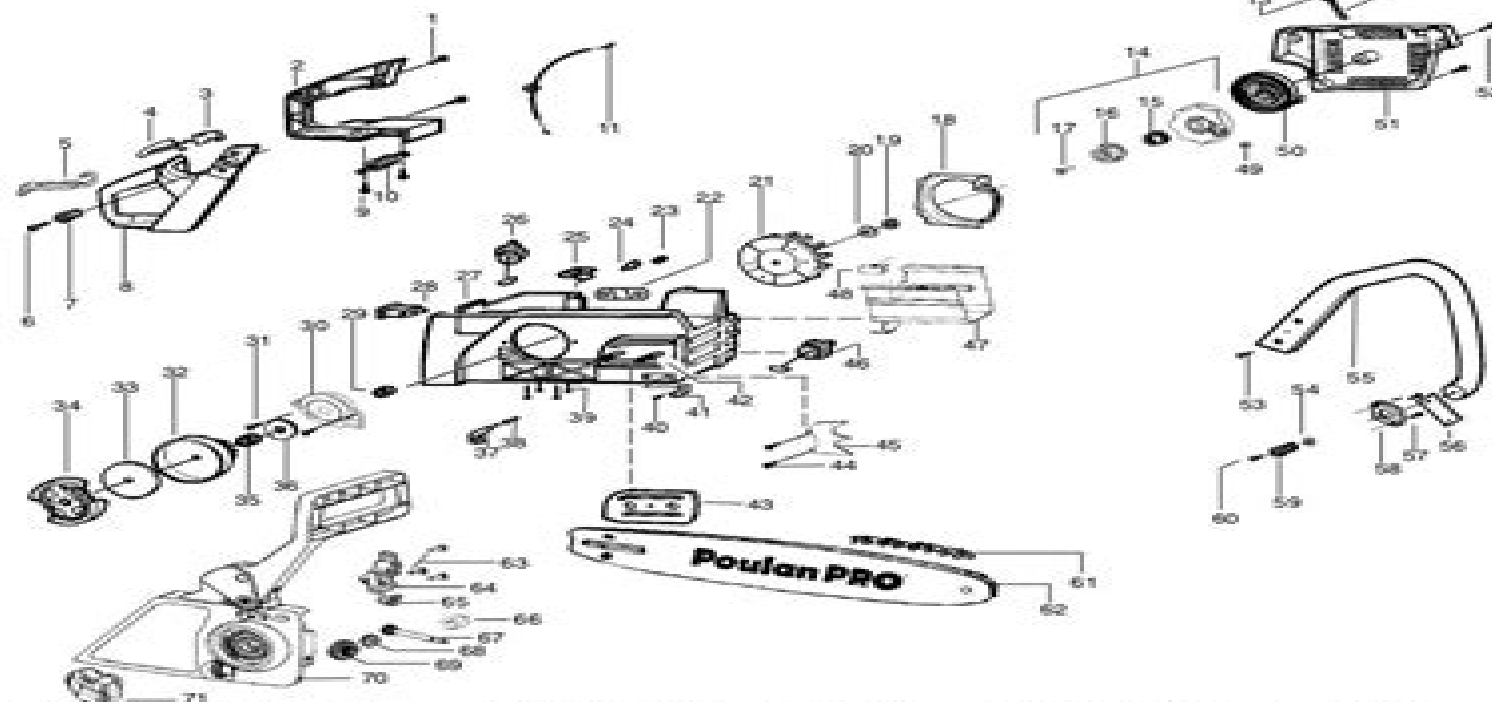


PARTS LIST NO. 545103450		Poulan PRO PARTS LIST	Chain Saw Model(s): PP-4620A/X
DATE: 04/15/09	Replaces: 54503010 - 7/03/06		PAGE NO. 2

Note: Illustration may differ from actual model due to design changes.



Ref.	Part No.	Description	Ref.	Part No.	Description	Ref.	Part No.	Description
1.	530015957	Screw	26.	530010845	Assy-Fuel Cap	49.	530016080	Screw
2.	530071272	Cover-Handle (kit)	27.	530037551	Assy-Chassis	50.	530042095	Spring-Recoil
3.	530036097	Trigger-Throttle			(Incl. 22,23,24,25,42,46	51.	545008803	Housing-Fan
4.	530036098	Lockout-Throttle			Fuel line and filter)	52.	530015892	Screw
5.	530053443	Lever-Choke	28.	530069572	Kit-Switch	53.	530015814	Screw
6.	530015937	Screw	29.	530029833	Worm-Oil Pump	54.	530015158	Washer
7.	530038965	Spring-Isolator	30.	530014410	Assy-Oil Pump	55.	530038327	Handle-Front
8.	530071272	Handle-Rear (kit)	31.	530015896	Screw	56.	530036113	Front Strap-AV
9.	530015906	Screw	32.	530069342	Kit-Clutch Drum	57.	530015775	Screw-Plate
10.	530038966	Spring-Isolator			(Incl. #35)	58.	530036101	Plate-Isolator
11.	530038471	Assy-Throttle Cable	33.	530015611	Washer-Clutch	59.	530038985	Spring-Isolator
		(Incl. Grommet)	34.	530014949	Assy-Glutch	60.	530015984	Screw
12.	530069232	Kit-Rope	35.	530032119	Bearing-Clutch	61.	952051310	Chain"- 20"
13.	530056422	Handle-Starter	36.	530015907	Washer-Thrust	62.	545049001	Bar -20"
14.	530071966	Kit-Starter EPs	37.	545047801	Pickup-Oil	63.	530016064	Screw
		(Incl. 15,16,17,49)	38.	530069533	Filter-Oil	64.	545051201	Cover-Gear
15.	530021180	Spring-EPs	39.	530016037	Screw	65.	545060501	Pin-Bar Adjust
16.	530059677	Hub-Starter EPs	40.	530015814	Screw	66.	545149301	Clip-Spring
17.	530021179	Screw-Shoulder	41.	530029850	Chain Catcher	67.	545060401	Screw-Adjuster
18.	530047667	Plate-Baffle Fan	42.	530015877	Bar Bolt	68.	545060201	Gear-Idler
19.	530014134	Nut-Flywheel	43.	530047855	Bar Plate	69.	545060501	Wheel-Gear Adjust
20.	530015127	Washer-Flywheel	44.	530016439	Screw	70.	545081874	Kit-ChainBrake Assy.
21.	530059637	Assy-Flywheel	45.	530014381	Kit-Spike Assy(Incl#44)			(Incl. Knob)
22.	530036043	Plate-Support	46.	530057236	Assy-Oil Cap			Kit - EZT Knob
		(Bar Stud)	47.	530055406	Shield-Heat			
23.	530029868	Plug-Vent	48.	530055322	Cap-"U"			
24.	530026119	Valve-Check						
25.	530015902	Nut-Speed						

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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized Ca(OH)_2 when calcium is put in water reacts as follows $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant K is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of K

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web sep 18 2020 in this experiment we will titrate a saturated calcium hydroxide solution with a standardized hcl solution we will determine the hydroxide ion concentration for the saturated calcium hydroxide

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared Ca(OH)_2 solution with a standard hydrochloric acid solution

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

solved chem 182 determining the ksp of calcium hydroxide chegg - Feb 18 2022

web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

ksp for calcium hydroxide university of manitoba - Dec 31 2022

web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

ksp calcium hydroxide chem132 determination of the ksp of calcium - Jun 24 2022

web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the

solution Ca(OH)_2 à Ca^{2+} $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide Ksp using the concentrations of calcium ions and hydroxide ions

solubility of calcium hydroxide wpmu dev - Oct 29 2022

web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the Ksp of a slightly soluble basic salt and how solubility changes with temperature Ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web question experiment 8 determining Ksp for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 HCl molarity 0.050 M volume of filtered Ca(OH)_2 solution mL 5.00 5.00 3.71 13.75 equivalence point volume mL OH in saturated Ca(OH)_2 M Ca in saturated Ca(OH)_2 M Ksp of Ca(OH)_2 average Ksp of Ca(OH)_2

chemistry 1210 lab 11 experiment 11 determining the Ksp of calcium - Sep 08 2023

web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be Ksp $\text{Ca}^{2+} \text{OH}^-$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18.1 solubility product constant Ksp chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as Ca_3PO_4 is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $\text{K}_{\text{Ca}_2\text{PO}_3\text{PO}_4} = \text{K}_{\text{Ca}_3\text{PO}_4} = \text{K}_{\text{Ca}_2\text{PO}_3\text{PO}_4}$

post lab report 8 determining the Ksp of calcium hydroxide - Jul 06 2023

web lab report 8 determining of calcium hydroxide the purpose of this was to titrate a saturated Ca(OH)_2 solution with a standard HCl solution and find the Ksp or the OH concentration was found and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochloric acid

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web question chem 182 experiment 8 determining the Ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of Ca(OH)_2 is represented in equation form lution of Ca on s nronic solid that is sparingly soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web plasma membrane homework symbiosis practice worksheet benzoin condensation grade a determining the ksp of calcium hydroxide after reading the lab complete items a b c and d title purpose chemicals and equipment and summary of procedure as described on page 10 of exp 1 on an 8 1 2 x 11 sheet of paper

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web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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web calculate the ksp for calcium hydroxide $\text{Ksp Ca(OH)}_2 = 5.0 \times 10^{-6}$ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5.0×10^{-6}

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