



Sea Ray Guide

Yan Bai



Sea Ray Guide:

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web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0

38 grams 0 02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric

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web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at stp $V_1/n_1 = V_2/n_2$ $0.0896 \text{ l} / n_1 = 22.4 \text{ l} / 1 \text{ mole}$ $n_1 = 0.0040 \text{ moles}$ 12 experimental molar mass g/mole of butane 0.23 g / 0.0040 moles = 58 g/mole 13 the molecular formula of butane is C_4H_{10}

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