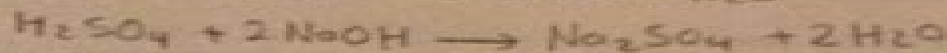
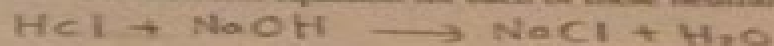


## Post-Laboratory Questions

(Use the spaces provided for the answers and additional paper if necessary.)

1. Obtain from your classmates or from your laboratory instructor, the molar enthalpies of neutralization of hydrochloric acid, sulfuric acid, and phosphoric acid when each reacts with sodium hydroxide.

(i) Write a chemical equation for each of these neutralization reactions.



(ii) Compare the molar enthalpies of neutralization of hydrochloric acid, sulfuric acid, and phosphoric acid. Discuss the similarities and differences of these data.

The enthalpy of neutralization is measured as per the number of moles of water formed. when one mole of water is formed the enthalpy of neutralization is  $-57.7 \text{ kJ}$ . In case of  $\text{HCl}$  and  $\text{H}_2\text{SO}_4$  one, two, and three moles of water are formed respectively. Hence, enthalpy of neutralization will be  $-57.7 \text{ kJ}$ ,  $2 \times -57.7 \text{ kJ}$ , and  $3 \times -57.7 \text{ kJ}$ , and

2. When a neutralization reaction was carried out using  $100.0 \text{ mL}$  of  $0.7890 \text{ M}$   $\text{NH}_3$ , water and  $100.0 \text{ mL}$  of  $0.7940 \text{ M}$  acetic acid,  $\Delta T$  was found to be  $4.76^\circ\text{C}$ . The specific heat of the reaction mixture was  $4.104 \text{ J g}^{-1} \text{ K}^{-1}$  and its density was  $1.03 \text{ g mL}^{-1}$ . The calorimeter constant was  $3.36 \text{ J K}^{-1}$ .

(i) Calculate  $\Delta H_{\text{mixture}}$  for the reaction of  $\text{NH}_3$  and acetic acid.

$$\text{mass of solution} = \text{total volume} \times \text{density}$$

$$= 200 \text{ mL} \times 1.03 \text{ g/mL} = 206.0 \text{ g}$$

$$\text{Heat released} = \text{mass} \times \text{specific heat capacity} \times \text{change in temp}$$

$$= 206.0 \times 4.104 \times 4.76$$

$$= 40.24 \text{ kJ}$$

$$\frac{40.24 \text{ kJ}}{\text{mixture}}$$

- (ii) At the end of the experiment, it was discovered that the thermometer had not been calibrated. When it was calibrated, it was found that the thermometer read  $0.50^\circ\text{C}$  low. What effect would this thermometer reading have on the reported  $\Delta H_{\text{mixture}}$  calculated in (i) above?

The uncalibrated thermometer reads  $0.50^\circ\text{C}$  lower than actual, therefore the actual temperature is

$$4.76 - 0.5 = 4.26^\circ\text{C}$$

$$\text{Heat released} = 206.0 \times 4.104 \times 4.26$$

$$= 3603.3 \text{ J}$$

$$= 36.03 \text{ kJ}$$

# Post Lab Question Answers To Chem Fax Labs

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