

Quiz: Oxidation of Elements

Instructions: Choose the best answer for each question. Each question assesses your understanding of the oxidation of elements and related concepts.

1. What is oxidation in terms of electron transfer?

- A) Gain of electrons
- B) Loss of electrons
- C) Gain of protons
- D) Loss of neutrons

2. Which of the following reactions represents oxidation?

- A) $2H_2 + O_2 \rightarrow 2H_2O$
- B) $Fe^{2+} + 2e^- \rightarrow Fe$
- C) $2Na + Cl_2 \rightarrow 2NaCl$
- D) $Cu + 2Ag^+ \rightarrow Cu^{2+} + 2Ag$

3. In a redox reaction, what does the term "reducing agent" refer to?

- A) The substance that gets oxidized
- B) The substance that gets reduced
- C) The substance that remains unchanged
- D) The substance that produces energy

4. Which element is most likely to be oxidized in the reaction: $Zn + Cu^{2+} \rightarrow Zn^{2+} + Cu$?

- A) Zn
- B) Cu^{2+}
- C) Zn^{2+}
- D) Cu

5. What is the oxidation state of sulfur in H_2SO_4 ?

- A) +2
- B) +4
- C) +6
- D) 0

6. Which of the following statements is true regarding oxidation states?

- A) Oxidation states can only be positive.
- B) The oxidation state of oxygen is always -2.
- C) Oxidation states are always whole numbers.

Oxidation State Word Answer Key

JG Myers



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