

Group 15 Elements

N, P, As, Sb, Bi, Mc

Hydrides of Group 15 Elements

Halides of Group 15 Elements The valence shell electronic configuration is ns²np³

N and P - non-metals
As and Sb - metalloids
Bi and Mc (moscovium) - metals

Common oxidation states are -3, +3 and +5 NH₃ has higher melting and boiling points than PH₃ due to hydrogen bonding

NH₃ is a mild reducing agent and BiH₃ is the strongest reducing agent

Basicity order - NH₃ > PH₃ > AsH₃ > SbH₃ > BiH₃

Pentahalides are more covalent than trihalides

Nitrogen does not form pentahalide due to absence of the d-orbitals in its valence shell

P Block Elements Class 12 Full Chapter

Anthony S. Fauci

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