

# Nomenclature

## Naming & Writing Chemical Formulas

### GROUP STUDY GUIDE

#### Part 1: Naming & Writing Ionic Compounds

- A. Naming Monatomic Ions  
 1. Cations: positive charge  
 2. Anions: negative charge

- B. Naming Binary Ionic Compounds  
 1. Cation: positive charge  
 2. Anion: negative charge

**Monatomic Ions**

**Cations:** Positive charge. Formed by the loss of electrons. Example:  $\text{Na}^+$ ,  $\text{Ca}^{2+}$ .

**Anions:** Negative charge. Formed by the gain of electrons. Example:  $\text{Cl}^-$ ,  $\text{O}^{2-}$ .

**Binary Ionic Compounds:** Compounds composed of two elements, one a metal and one a nonmetal. Example:  $\text{NaCl}$ ,  $\text{CaO}$ .

**Naming:** The cation is named first, followed by the anion. Example:  $\text{NaCl}$  is named sodium chloride.

**Writing:** The formula is written by combining the symbols of the ions, with subscripts indicating the number of each ion. Example:  $\text{NaCl}$ ,  $\text{CaO}$ .

**Examples:**

- $\text{Na}^+$  (Sodium ion)
- $\text{Ca}^{2+}$  (Calcium ion)
- $\text{Cl}^-$  (Chloride ion)
- $\text{O}^{2-}$  (Oxide ion)
- $\text{NaCl}$  (Sodium chloride)
- $\text{CaO}$  (Calcium oxide)

**Binary Ionic Compounds**

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**Examples:**

- $\text{NaCl}$  (Sodium chloride)
- $\text{CaO}$  (Calcium oxide)
- $\text{MgS}$  (Magnesium sulfide)
- $\text{Al}_2\text{O}_3$  (Aluminum oxide)

#### Part 2: Naming & Writing Molecular Compounds

**Molecular Compounds:** Compounds composed of two or more nonmetals. Example:  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ .

**Naming:** The first element is named first, followed by the second element with a prefix. Example:  $\text{CO}_2$  is named carbon dioxide.

**Writing:** The formula is written by combining the symbols of the elements, with subscripts indicating the number of each element. Example:  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ .

**Examples:**

- $\text{CO}_2$  (Carbon dioxide)
- $\text{H}_2\text{O}$  (Water)
- $\text{CH}_4$  (Methane)
- $\text{N}_2$  (Nitrogen)

**Molecular Compounds**

**Naming:** The first element is named first, followed by the second element with a prefix. Example:  $\text{CO}_2$  is named carbon dioxide.

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- $\text{CH}_4$  (Methane)
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# Nomenclature Study Guide

**Emilie Sanchez**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

## **Nomenclature Study Guide:**

## Unveiling the Magic of Words: A Overview of "**Nomenclature Study Guide**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Nomenclature Study Guide**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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