

Name: Key

Class: \_\_\_\_\_

Date: \_\_\_\_\_

Stoichiometry Problem Set A

**Directions:** Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

**Part 1: Straight Stoichiometry Problems**

**Target Goal:** To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

- 1.) For the reaction  $2\text{HNO}_3 + 1\text{Mg(OH)}_2 \rightarrow 1\text{Mg(NO}_3)_2 + 2\text{H}_2\text{O}$ , how many grams of magnesium <sup>B</sup> nitrate are produced from 8.00 mol of nitric acid ( $\text{HNO}_3$ )? A

$$\text{MM} = 148.3\text{g}$$

$$\frac{8.00\text{mol HNO}_3}{1} \times \frac{1\text{mol Mg(NO}_3)_2}{2\text{mol HNO}_3} \times \frac{148.3\text{g Mg(NO}_3)_2}{1\text{mol Mg(NO}_3)_2} = 592.8\text{g Mg(NO}_3)_2$$

- 2.) For the reaction  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ , how many moles of potassium <sup>B</sup> chlorate are required to produce 160. g of oxygen? A

$$\text{MM} = 32.0\text{g}$$

$$\frac{160.\text{g O}_2}{1} \times \frac{1\text{mol O}_2}{32.0\text{g O}_2} \times \frac{2\text{mol KClO}_3}{3\text{mol O}_2} = 3.33\text{ mole KClO}_3$$

- 3.) For the reaction  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ , how many grams of water are produced from 5.22 liters of oxygen gas? <sup>B</sup> A

$$\text{MM} = 18.0\text{g}$$

$$\frac{5.22\text{L O}_2}{1} \times \frac{1\text{mol O}_2}{22.4\text{L O}_2} \times \frac{2\text{mol H}_2\text{O}}{1\text{mol O}_2} \times \frac{18.0\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 8.39\text{g H}_2\text{O}$$

- 4.) In the reaction  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ , how many grams of hydrogen <sup>B</sup> are produced if 1.33 grams of sodium react completely? A

$$\text{MM} = 23.0\text{g}$$

$$\text{MM} = 2.0\text{g}$$

$$\frac{1.33\text{g Na}}{1} \times \frac{1\text{mol Na}}{23.0\text{g Na}} \times \frac{1\text{mole H}_2}{2\text{mole Na}} \times \frac{2.0\text{g H}_2}{1\text{mole H}_2} = 0.0578\text{g H}_2$$

- 5.) In the reaction  $1\text{CO}_2 + 1\text{H}_2\text{O} \rightarrow 1\text{H}_2\text{CO}_3$ , how many moles of carbonic <sup>B</sup> acid can be produced from 97.2 g of  $\text{CO}_2$ ? A

$$\text{MM} = 44.0\text{g}$$

$$\frac{97.2\text{g CO}_2}{1} \times \frac{1\text{mole CO}_2}{44.0\text{g CO}_2} \times \frac{1\text{mole H}_2\text{CO}_3}{1\text{mole CO}_2} = 2.21\text{ moles H}_2\text{CO}_3$$

**Did you achieve our target goal?**

**Target Goal:** To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

1

Feeling VERY confused

2

Not sure that I get this

3

I think I am catching on

4

Doing great! Bring it on!

# Sample Problem Set Stoichiometry Chemistry

**RS Peters**



## Sample Problem Set Stoichiometry Chemistry:

**A Visual Analogy Guide to Chemistry, 2e** Paul A Krieger, 2018-02-01 A Visual Analogy Guide to Chemistry is the latest in the innovative and widely used series of books by Paul Krieger This study guide delivers a big picture view of difficult concepts and effective study tools to help students learn and understand the details of general organic and biochemistry topics A Visual Analogy Guide to Chemistry is a worthwhile investment for any introductory chemistry student

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Chemical Education: Towards Research-based Practice J.K. Gilbert, Onno de Jong, Rosária Justi, David F. Treagust, Jan H. van Driel, 2006-03-11 Chemical education is essential to everybody because it deals with ideas that play major roles in personal social and economic decisions This book is based on three principles that all aspects of chemical education should be associated with research that the development of opportunities for chemical education should be both a continuous process and be linked to research and that the professional development of all those associated with chemical education should make extensive and diverse use of that research It is intended for pre service and practising chemistry teachers and lecturers chemistry teacher educators chemical education researchers the designers and managers of formal chemical curricula informal chemical educators authors of textbooks and curriculum support materials practising chemists and chemical technologists It addresses the relation between chemistry and chemical education curricula for chemical education teaching and learning about chemical compounds and chemical change the development of teachers the development of chemical education as a field of enquiry This is mainly done in respect of the full range of formal education contexts schools universities vocational colleges but also in respect of informal education contexts books science centres and museums

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