



NATIONAL
EXAM CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P1
(CHEMISTRY),
EXEMPLAR

MARKS
100



The question paper consists of 10 pages, including this cover page.
There are 10 questions in this question paper.

QUESTION 3

PHYSICAL SCIENCES PAPER 2 INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

QUESTION 3 MARKS AVAILABLE 10

QUESTION 3 TIME ALLOWED 30 minutes

- 3.1 The following table shows the interaction energy observed during an intermolecular force experiment for halogen halide dihalides. Use this table and the table on the back page to answer the following questions.

COMPOUND	INTERACTION ENERGY (KJ/mol)	PERCENTAGE (%)
F_2	-100	100
Cl_2	-100	0

- 3.1.1 Calculate the percentage:
a) of F_2 placed on compound Cl_2 and Cl_2 placed on F_2 .
b) from the type of intermolecular force that each compound exhibits. State F_2 and Cl_2 in the order named.
c) which child in the classroom did the boy with a different colour hair dye probably have?
d) the number of pairs present.
e) which substance has the highest dipole-dipole polarisability?
f) the reason for the student in referring to the name of the substance in 3.1.1, rather than the name of a similar substance that does not contain fluorine or chlorine, as having higher binding forces.
g) Explain why the binding force of chlorine is stronger than that of fluorine. Give two reasons for this.
h) State which child in the class has a stronger dipole-dipole interaction force.



- 3.1.2 In which liquid will each of the following dissolve?
a) Copper(II) sulphate
b) Sodium hydroxide
c) Zinc sulphide



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Physical Science Paper 2 Intermolecular Forces

Investigation

J.S. Rowlinson

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Study Guide for CTET Paper 2 (Class 6 - 8 Teachers) Mathematics/ Science with Past Questions Disha Experts,2020-02-04 Proceedings of the Royal Society. Section A, Mathematical and Physical Science Royal Society (Great Britain),1908

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Scientific and Technical Aerospace Reports ,1995 Nuclear Science Abstracts ,1971 Sir James Dewar, 1842-1923 J.S. Rowlinson,2016-03-09 Sir James Dewar was a major figure in British chemistry for around 40 years He held the posts of Jacksonian Professor of Natural Philosophy at Cambridge 1875 1923 and Fullerian Professor of Chemistry at the Royal Institution 1877 1923 and is remembered principally for his efforts to liquefy hydrogen successfully in the field that would come to be known as cryogenics His experiments in this field led him to develop the vacuum flask now more commonly known as the thermos and in 1898 he was the first person to successfully liquefy hydrogen A man of many interests he was also with Frederick Abel the inventor of explosive cordite an achievement that involved him in a major legal battle with Alfred Nobel Indeed Dewar s career saw him involved in a number of public quarrels with fellow scientists he was a fierce and sometimes unscrupulous defender of his rights and his claims to priority in a way that throws much light on the scientific spirit and practice of his day This the first scholarly biography of Dewar seeks to resurrect and reinterpret a man who was a

giant of his time but is now sadly overlooked In so doing the book will shed much new light on the scientific culture of the late nineteenth and early twentieth centuries and the development of the field of chemistry in Britain *U.S. Government Research & Development Reports* ,1970-04 [Resources in Education](#) ,1972-10 **Journal of Research of the National Bureau of Standards** ,1963 *High-Performance Computing Infrastructure for South East Europe's Research*

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