

## FINAL EXAMANITION IN PHYSICAL SCIENCE

**I. Multiple Choices:** For each of the following items, choose the BEST answer. On the answer sheet provided, shade the bubble that corresponds to the letter of your choice.

- The chemical formula of caffeine is  $C_8H_{10}N_4O_2$ . What is its molecular mass?  
A. 165 g/mol      B. 172 g/mol      C. 194 g/mol      D. 208 g/mol
- People usually use hydrogen peroxide ( $H_2O_2$ ; molar mass = 34.02 g/mol) to clean their wounds. If Ricardo Delasay used 1.0 g of  $H_2O_2$  to clean his gunshot wound, how many moles of  $H_2O_2$  did he use?  
A. 0.035 mol      B. 0.030 mol      C. 0.029 mol      D. 0.0025 mol
- Methane ( $CH_4$ ; molar mass = 16.04 g/mol) is one of the greenhouses gases. If 32.08 g of  $CH_4$  was released in the atmosphere, how many molecules of  $CH_4$  were added in the atmosphere?  
A.  $1.20 \times 10^{26}$       B.  $1.20 \times 10^{23}$       C.  $6.02 \times 10^{26}$       D.  $6.02 \times 10^{23}$

For numbers 4 and 5, consider the problem below:

- Iron (III) oxide reacts with carbon monoxide to form iron and carbon dioxide. Which is the balanced chemical equation that represents the reaction?  
A.  $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$       C.  $2Fe_2O_3 + 3CO \rightarrow 4Fe + 3CO_2$   
B.  $Fe_2O_3 + CO \rightarrow FeCO_2$       D.  $Fe_2O_3 + 3CO \rightarrow Fe + 3CO_2$
- What mass of iron (III) oxide is needed to produce 5.6 kg of iron?  
A. 4.0 kg      B. 8.0 kg      C. 12 kg      D. 16 kg

For numbers 6-8, consider the problem below:

Ammonium carbonate decomposes to form ammonia gas, carbon dioxide and water. The reaction is represented by the equation below:



- How many moles of ammonia are formed for every mole of the reactant used?  
A. 1 mole      B. 2 moles      C. 2.5 moles      D. 3 moles
- How many moles of the reactant is required to produce 22 grams of  $CO_2$ ?  
A. 0.50 mole      B. 1.0 mole      C. 1.5 mole      D. 2.0 moles
- How many grams of  $(NH_4)_2CO_3$  is required to produce 1.5 moles of water?  
A. 49 g      B. 96 g      C. 144 g      D. 192 g

For numbers 9 and 10, consider the problem below:

45.9 g of  $K_2CO_3$  was reacted with an excess of  $HCl$ , yielding 48.25 g of  $KCl$ .



- What is the theoretical yield?  
A. 46.3 g      B. 48.0 g      C. 48.3 g      D. 49.5 g
- What is the percent yield?  
A. 92.5 %      B. 93.5 %      C. 97.0 %      D. 97.5 %
- Which of the following will cause a soap to reduce its cleaning activity?  
A. Alkyl hydrogen sulfate      C. presence of surfactants  
B. Hard water      D. soft water
- Which of the following chemicals are linked to breast cancer in women?  
A. Parabens      B. PEG      C. Paraffin wax      D. BHA
- Which of the following disinfectants has a fungicidal property and is a potential carcinogen?

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