

POGIL 2 Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Diagram 1 on the next page to complete the following table.

	Metal or Nonmetal	Number of protons in the atom	Number of protons in the ion	Number of neutrons in the atom	Number of neutrons in the ion	Number of electrons in the atom	Number of electrons in the ion	Charge on the ion
Lithium	metal	3	3	4	4	3	2	1+
Magnesium	<i>Metal</i>	12	12	12	12	12	10	2+
Aluminum	<i>Metalloid</i>	13	13	14	14	13	10	3+
Fluorine	<i>Non metal</i>	9	9	10	10	9	10	1-
Oxygen	nonmetal	8	8	8	8	8	10	2-
Nitrogen	<i>Non metal</i>	7	7	7	7	7	10	3-

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

A neutral atom has the same number of protons and electrons (no charge). An ion has a charge (# of protons does not equal # of electrons).

3. Examine the isotope symbols in Diagram 1.

- a. Where is the ion charge located in the isotope symbol?

Top right corner

- b. Is a charge indicated on the neutral atoms? If yes, where is it located?

No

4. Which subatomic particle carries a positive charge?

Protons are positive

5. Which subatomic particle carries a negative charge?

Electrons are negative

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Diagram 1.

The charge of an ion = # of protons - # of electrons

Polyatomic Ions Pogil Packet Answers

Aharon Loewenschuss, Y. Marcus



Polyatomic Ions Pogil Packet Answers:

Polyatomic Ions for Surface Analysis and Modification Erick Ryan Fuoco, 2003 Investigations Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry Sally M. McIntyre, 2010 **The Particulate Nature of Polyatomic Ions**, 2007 *High Resolution Studies of the Origins of Polyatomic Ions in Inductively Coupled Plasma-Mass Spectrometry* Jill Wisniewski Ferguson, 2006

The inductively coupled plasma ICP is an atmospheric pressure ionization source. Traditionally the plasma is sampled via a sampler cone. A supersonic jet develops behind the sampler and this region is pumped down to a pressure of approximately one Torr. A skimmer cone is located inside this zone of silence to transmit ions into the mass spectrometer. The position of the sampler and skimmer cones relative to the initial radiation and normal analytical zones of the plasma is key to optimizing the useful analytical signal.

1 The ICP both atomizes and ionizes the sample. Polyatomic ions form through ion molecule interactions either in the ICP or during ion extraction.

1 Common polyatomic ions that inhibit analysis include metal oxides. MO sup adducts with argon, the gas most commonly used to make up the plasma, and hydride species. While high resolution devices can separate many analytes from common interferences, this is done at great cost in ion transmission efficiency: a loss of 99% when using high versus low resolution on the same instrument.

2 Simple quadrupole devices, which make up the bulk of ICP MS instruments in existence, do not present this option. Therefore, if the source of polyatomic interferences can be determined and then manipulated, this could potentially improve the figures of merit on all ICP MS devices, not just the high resolution devices often utilized to study polyatomic interferences.

Fragmentation of Diatomic and Polyatomic Ions in the Gas Phase Christopher John Proctor, 1981

Investigations Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry, 2010

An inductively coupled plasma mass spectrometer (ICP MS) is an elemental analytical instrument capable of determining nearly all elements in the periodic table at limits of detection in the parts per quadrillion and with a linear analytical range over 8-10 orders of magnitude. Three concentric quartz tubes make up the plasma torch. Argon gas is spiraled through the outer tube and generates the plasma, powered by a looped load coil operating at 27.1 or 40.6 MHz. The argon flow of the middle channel is used to keep the plasma above the innermost tube, through which solid or aqueous sample is carried in a third argon stream. A sample is progressively desolvated, atomized, and ionized. The torch is operated at atmospheric pressure. To reach the reduced pressures of mass spectrometers, ions are extracted through a series of two approximately one millimeter wide circular apertures set in water-cooled metal cones. The space between the cones is evacuated to approximately one torr. The space behind the second cone is pumped down to or near to the pressure needed for the mass spectrometer. MS The first cone, called the sampler, is placed directly in the plasma plume and its position is adjusted to the point where atomic ions are most abundant. The hot plasma gas expands through the sampler orifice and in this expansion is placed the second cone, called the skimmer. After the skimmer, traditional MS designs are employed, i.e., quadrupoles, magnetic sectors, time of flight.

ICP MS is the leading trace element analysis technique. One of its weaknesses are polyatomic ions. This dissertation has added to the fundamental understanding of some of these polyatomic ions, their origins and behavior. Although mainly continuing the work of others, certain novel approaches have been introduced here. Chapter 2 includes the first reported efforts to include high temperature corrections to the partition functions of the polyatomic ions in ICP MS. This and other objections to preceding papers in this area were addressed. Errors in the measured T_{gas} values were found for given errors in the experimental and spectroscopic values. The ionization energy of the neutral polyatomic ion was included in calculations to prove the validity of ignoring more complicated equilibria. Work was begun on the question of agreement between kinetics of the plasma and interface and the increase and depletion seen in certain polyatomic ions. This dissertation was also the first to report day to day ranges for T_{gas} values and to use a statistical test to compare different operating conditions. This will help guide comparisons of previous and future work. Chapter 4 was the first attempt to include the excited electronic state 2 in the partition function of ArO as well as the first to address the different dissociation products of the ground and first electronic levels of ArO. Chapter 5 reports an interesting source of memory in ICP MS that could affect mathematical corrections for polyatomic ions. For future work on these topics, I suggest the following experiments and investigations. Clearly not an extensive list, they are instead the first topics curiosity brings to mind.

1. Measurement of T_{gas} values when using the flow injection technique of Appendix B. It was believed that there was a fundamental difference in the plasma when the auto sampler was used versus a continuous injection. Is this reflected in T_{gas} values?
2. The work of Chapter 3 can be expanded and supplemented with more trials, new cone materials, i.e. copper, stainless steel, and more cone geometries. Some of this equipment is already present in the laboratory; others could be purchased or made.
3. T_{gas} values from Chapter 3 could be correlated with instrument pressures during the experiment. Pressures after the skimmer cone were recorded for many days but have yet to be collated with the measured T_{gas} values.
4. **Fragmentation of Diatomic and Polyatomic Ions in the Gas Phase** Christopher John Proctor, 1981
- Unimolecular and Collision-induced Dissociation Study of Polyatomic Ions at High Collision Energy** Xuedong Zhou, 2001
- Standard Thermodynamic Functions of Gaseous Polyatomic Ions at 100-1000 K** Aharon Loewenschuss, Y. Marcus, 1987
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Table of Contents Polyatomic Ions Pogil Packet Answers

1. Understanding the eBook Polyatomic Ions Pogil Packet Answers
 - The Rise of Digital Reading Polyatomic Ions Pogil Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Polyatomic Ions Pogil Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Polyatomic Ions Pogil Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Polyatomic Ions Pogil Packet Answers
 - Personalized Recommendations
 - Polyatomic Ions Pogil Packet Answers User Reviews and Ratings
 - Polyatomic Ions Pogil Packet Answers and Bestseller Lists
5. Accessing Polyatomic Ions Pogil Packet Answers Free and Paid eBooks
 - Polyatomic Ions Pogil Packet Answers Public Domain eBooks
 - Polyatomic Ions Pogil Packet Answers eBook Subscription Services
 - Polyatomic Ions Pogil Packet Answers Budget-Friendly Options

6. Navigating Polyatomic Ions Pogil Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Polyatomic Ions Pogil Packet Answers Compatibility with Devices
 - Polyatomic Ions Pogil Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Polyatomic Ions Pogil Packet Answers
 - Highlighting and Note-Taking Polyatomic Ions Pogil Packet Answers
 - Interactive Elements Polyatomic Ions Pogil Packet Answers
8. Staying Engaged with Polyatomic Ions Pogil Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Polyatomic Ions Pogil Packet Answers
9. Balancing eBooks and Physical Books Polyatomic Ions Pogil Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Polyatomic Ions Pogil Packet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Polyatomic Ions Pogil Packet Answers
 - Setting Reading Goals Polyatomic Ions Pogil Packet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Polyatomic Ions Pogil Packet Answers
 - Fact-Checking eBook Content of Polyatomic Ions Pogil Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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