

Blackline Master 5.8

Ionic Compounds: Names and Formulas Worksheet

1. Write the formulas for the following compounds.

(a) magnesium oxide	<u>MgO</u>
(b) sodium fluoride	<u>NaF</u>
(c) aluminum nitride	<u>AlN</u>
(d) potassium sulfide	<u>K₂S</u>
(e) lithium iodide	<u>LiI</u>
(f) calcium bromide	<u>CaBr₂</u>
(g) beryllium oxide	<u>BeO</u>
(h) nickel ^(II) chloride	<u>NiCl₂</u>
(i) magnesium nitride	<u>Mg₃N₂</u>
(j) aluminum sulfide	<u>Al₂S₃</u>

(k) copper(I) bromide	<u>CuBr</u>
(l) tin(II) iodide	<u>SnI₂</u>
(m) iron(III) chloride	<u>FeCl₃</u>
(n) calcium phosphide	<u>Ca₃P₂</u>
(o) lead(II) oxide	<u>PbO</u>
(p) lead(IV) fluoride	<u>PbF₄</u>
(q) tin(IV) bromide	<u>SnBr₄</u>
(r) copper(II) sulfide	<u>CuS</u>
(s) iron(II) oxide	<u>FeO</u>
(t) calcium nitride	<u>Ca₃N₂</u>

2. Write the names for the following compounds.

(a) Li ₂ O	<u>lithium oxide</u>
(b) AlCl ₃	<u>aluminum chloride</u>
(c) MgS	<u>magnesium sulfide</u>
(d) CaO	<u>calcium oxide</u>
(e) KBr	<u>potassium bromide</u>
(f) BeF	<u>beryllium fluoride</u>
(g) Na ₃ N	<u>sodium nitride</u>
(h) Al ₂ O ₃	<u>aluminum oxide</u>
(i) CuCl ₂	<u>copper(II) chloride</u>
(j) FeBr ₃	<u>iron(III) bromide</u>

(k) PbS	<u>lead (II) sulfide</u>
(l) SnO ₂	<u>tin (IV) oxide</u>
(m) Na ₂ S	<u>sodium sulfide</u>
(n) Mg ₃ P ₂	<u>magnesium phosphide</u>
(o) NiO	<u>nickel (II) oxide</u>
(p) CuI	<u>copper (I) iodide</u>
(q) PbCl ₂	<u>lead (II) chloride</u>
(r) FeP	<u>iron (III) phosphide</u>
(s) CaF ₂	<u>calcium fluoride</u>
(t) K ₃ P	<u>potassium phosphide</u>

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ML Morrison



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