		gerangskik sis om die hoeke tussen hulle te maksimeer. Strukturele elektronpare is bindingspare plus alleenpare.  OF  Valensdop-elektronpaarafstotingsmodel is 'n model om die vorms van molekules te voorspels waarin die elektronpaar vanaf die buitenste dops  van 'n verwysingsatoom rondom hierdie atoom gerangskik is om die afstoting  tussen hulle te minimaliseer.	
	1.2.4	The molecule is a tetrahedron <sup>™</sup> Die molekule is 'n tetraëder <sup>™</sup> 1 mark/punt	LO2
	1.2.5	Name trigonal pyramid  Noem trigonale piramidale  3 marks/punte	LO4
1.3	1.3.1	The electron pair is equally shared between the two CI atoms and the resulting Cl <sub>2</sub> molecule has no degree of polarity. Die elektronpaar word ewe veel gedeel tussen die twee CI-atome en die resulterende CI2-molekule het geen mate van polariteit nie. 2 marks/punte	LO2
	1.3.2	The tendency of an atom to pull the bonded pair of electrons towards itself. of Die neiging of van 'n atoom om die gebinde elektronpare na homself te trek of a marks/punte	LO2
	1.3.3	Electronegativity of Na = 6 Electronegativity of Cl = 3  Difference is 3   With a difference of greater than 2.1 electron transfer will take place and the bond elis ionic. (value of 2.1 does not need to be given for this mark) Elektronegatiwiteit van Na = 6 Elektronegatiwiteit van Cl = 3  Verskil is 3   Met 'n verskil van meer as 2,1 sal elektronoordrag plaasvind en die binding is ionies. (waarde van 2.1 hoef nie vir hierdie punt gegee te word nie) 4 marks/punte	LO2

# **Physical Science Control Test February 2014**

**G Psacharopoulos** 

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## **Table of Contents Physical Science Control Test February 2014**

- 1. Understanding the eBook Physical Science Control Test February 2014
  - The Rise of Digital Reading Physical Science Control Test February 2014
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Control Test February 2014
  - Exploring Different Genres
  - $\circ\,$  Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Control Test February 2014
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Control Test February 2014
  - Personalized Recommendations
  - Physical Science Control Test February 2014 User Reviews and Ratings
  - Physical Science Control Test February 2014 and Bestseller Lists
- 5. Accessing Physical Science Control Test February 2014 Free and Paid eBooks
  - Physical Science Control Test February 2014 Public Domain eBooks
  - Physical Science Control Test February 2014 eBook Subscription Services
  - Physical Science Control Test February 2014 Budget-Friendly Options
- 6. Navigating Physical Science Control Test February 2014 eBook Formats
  - o ePub, PDF, MOBI, and More
  - Physical Science Control Test February 2014 Compatibility with Devices
  - Physical Science Control Test February 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Control Test February 2014
  - Highlighting and Note-Taking Physical Science Control Test February 2014
  - Interactive Elements Physical Science Control Test February 2014
- 8. Staying Engaged with Physical Science Control Test February 2014

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Physical Science Control Test February 2014
- 9. Balancing eBooks and Physical Books Physical Science Control Test February 2014
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Control Test February 2014
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Control Test February 2014
  - Setting Reading Goals Physical Science Control Test February 2014
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Control Test February 2014
  - Fact-Checking eBook Content of Physical Science Control Test February 2014
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - $\circ \ \ Integration \ of \ Multimedia \ Elements$
  - Interactive and Gamified eBooks

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