

Name \_\_\_\_\_

# Physical Science Chemical Bonding Study Guide

## A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 2-? \_\_\_\_\_ (anion)  
which cation is always a 2+? \_\_\_\_\_ (cation)
2. Is it placed in the halogen group has \_\_\_\_\_ valence electrons, and needs to \_\_\_\_\_ 1 more to have 8. This is because it has \_\_\_\_\_ valence.
3. Is it \_\_\_\_\_ bond to the adjacent elements oppositely charged ions.  
PCl<sub>3</sub> is covalent bond is \_\_\_\_\_
4. Is it considered ionic, atomic, \_\_\_\_\_, molecular, defined by an  
ionic bond, atomic, \_\_\_\_\_, molecular.
5. 2-bonded group: Which side bond has shared completely by a covalent bond is to  
\_\_\_\_\_ and others group's shared equally to the  
\_\_\_\_\_ group. If it is covalent, its bond, polar covalent, or ionic bond is determined by  
electronegativity difference \_\_\_\_\_
6. Is it atomic or ionic bond has \_\_\_\_\_ valence electrons. This family includes all the  
periodic table that all have 8 valence electrons, is called \_\_\_\_\_
7. Atomic ions gain, lose, or share \_\_\_\_\_ by becoming more  
stable.

## B. Formulas for Ionic Compounds - Put in the right order

Write the correct electronegativity difference is	Alkali metals	Transition	Post-transition	Nonmetals

3. Write in the chemical formula for each of the following ions:  
compounds of

Write your own ionic compounds, and use arrows to show the transfer of  
electrons. Use your periodic table to find their oxidation numbers. They will combine  
so that their oxidation numbers cancel out to zero. Use the criss cross method if their  
oxidation numbers do not cancel.

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**John L. Rudolph**



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