
Nocti Study Guide

Leg - Answer- The proper method for lifting is to keep the back straight, spread the feet comfortably and lift with _____muscles.

Shut off the power - Answer- If a light catches on fire, the first step is to

Touch the lamp only with protected gloves - Answer- When replacing a halogen lamp in a lighting instrument

It potentially saves time and money - Answer- Why is a Production schedule important?

Treatment - Answer- In planning the shots and audio for a television commercial, the Director oversees the development of a/an

Streaming - Answer- When showing video over the Internet, the sender must request _____rights from the publishers for any non-original material used.

Auditions - Answer- Which of the following is used to test the abilities of the Talent?

Quick-Release Plate - Answer- What Tripod mechanism facilitates easy and accurate mounting of the camcorder?

OTS - Answer- What is the script notation for framing the subject with a person in the foreground?

Technical Director - Answer- Which of the following personnel operates the switcher during a studio production?

Put them away last to give them time to cool off - Answer- When shooting video on location, to make sure the lights have cooled, the operator should

Rating Information - Answer- Nielsen provides TV and radio stations with

Treatment - Answer- A brief narrative description for a Television program is called a/an

Where cables run - Answer- What technical requirements should be considered during a site survey or location scout?

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L Darling-Hammond



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + 2 \text{ H}_2(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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web chemistry 120 hess s law worksheet calculate Δh for the reaction $\text{C}_2\text{H}_4(\text{g}) + 2 \text{ C}_2\text{H}_6(\text{g}) \rightarrow 3 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ}$ mole $\text{C}_2\text{H}_6(\text{g})$ $7 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ}$ mole $\text{H}_2(\text{g})$ $1 \text{ O}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ}$ mole

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web hess s law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall process is the sum of the Δh values for each step to use hess s law two principles must be understood one if an equation is reversed the sign of the Δh value is also reversed two if an equation is multiplied by a

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web hess s law worksheet answers 1 calculate h for the reaction $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$ from the following data $\text{C}_2\text{H}_4 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2\text{O l} \text{ h } 1411 \text{ kj}$ $\text{C}_2\text{H}_6 \text{ g} + 3\frac{1}{2} \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O l} \text{ h } 1560 \text{ kj}$

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O h 55 8 kj as shown note that the enthalpy change $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta h_{\text{products}}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$ $\Delta H = -44.5 \text{ kJ/mol}$ 2 $\text{NaOH (s)} + \text{HCl (aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{Cl}^- \text{(aq)} + \text{H}_2\text{O (l)}$ $\Delta H = -57.3 \text{ kJ/mol}$ 3