

Lab POGIL: Metal Reactivity – Introduction & Pre-Lab

Discovery Question: *How can trends in metal reactivity be used to predict the products of chemical reactions?*

Objectives:

- Develop a procedure to test the reactivity of metals
- List metals in order of reactivity based on the results of your procedure
- Predict the products of chemical reactions by considering the relative reactivity of the reactants

You have been provided with vials containing 3 types of metals (A, B, C). You also have vials filled with solutions containing cations of each metal (A^{+} , B^{+} , C^{+}), and an additional vial containing a solution of cation D^{+} . You will also have (6-well x 4-well) plates available for small-scale reactions of metals and ionic solutions.

Elemental Metals	Solutions Containing Metal Cations
A	A (contains A^{+})
B	B (contains B^{+})
C	C (contains C^{+})
	D (contains D^{+})

Remember to wear goggles at all times. Avoid skin contact with all solutions. If contact occurs, rinse with water immediately. Solution D will stain skin and clothes.

Pre-lab Requirements:

- Describe a method for predicting charges of ions formed from single atoms of an element using the periodic table.
- Use a periodic table to predict charges for ions formed from each element listed below:

Element	Ion	Element	Ion	Element	Ion
Lithium (Li)	_____	Beryllium (Be)	_____	Copper (Cu)	____/____
Sodium (Na)	_____	Magnesium (Mg)	_____	Zinc (Zn)	_____
Potassium (K)	_____	Calcium (Ca)	_____	Silver (Ag)	_____
Nitrogen (N)	_____	Oxygen (O)	_____	Fluorine (F)	_____
Phosphorous (P)	_____	Sulfur (S)	_____	Chlorine (Cl)	_____
		Selenium (Se)	_____	Bromine (Br)	_____

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