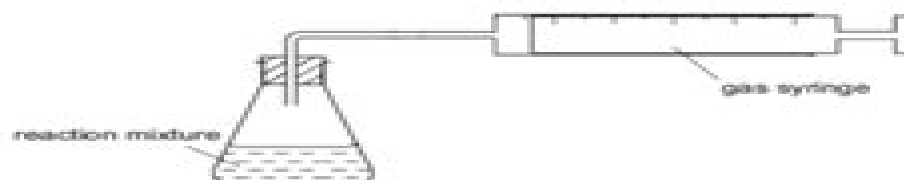


Reaction rates

0620 Cambridge IGCSE Chemistry Specimen Papers (2016)

Paper 1 Question 15

15. The apparatus shown can be used to measure the rate of some chemical reactions.



For which two reactions would this apparatus be suitable?

- | | |
|------------|--|
| reaction 1 | $\text{AgNO}_3(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{HNO}_3(\text{aq})$ |
| reaction 2 | $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ |
| reaction 3 | $\text{MgO}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l})$ |
| reaction 4 | $\text{ZnCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{ZnCl}_2(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ |

- A 1 and 2 B 1 and 3 C 2 and 4 D 3 and 4

Paper 1 Question 16

16. A student investigates the rate of reaction between magnesium and excess sulfuric acid.

The volume of hydrogen given off in the reaction is measured over time.

The graph shows the results of two experiments, R and S.



Which change in conditions would cause the difference between R and S?

- A A catalyst is added in S.
 B The acid is more concentrated in R than in S.
 C The magnesium is less finely powdered in R than in S.
 D The temperature in R is lower than in S.

Reaction Rates Chemistry Packet Answers

J Spring



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Table of Contents Reaction Rates Chemistry Packet Answers

1. Understanding the eBook Reaction Rates Chemistry Packet Answers
 - The Rise of Digital Reading Reaction Rates Chemistry Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates Chemistry Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rates Chemistry Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates Chemistry Packet Answers
 - Personalized Recommendations
 - Reaction Rates Chemistry Packet Answers User Reviews and Ratings
 - Reaction Rates Chemistry Packet Answers and Bestseller Lists

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 - Adjustable Fonts and Text Sizes of Reaction Rates Chemistry Packet Answers
 - Highlighting and Note-Taking Reaction Rates Chemistry Packet Answers
 - Interactive Elements Reaction Rates Chemistry Packet Answers
8. Staying Engaged with Reaction Rates Chemistry Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rates Chemistry Packet Answers
9. Balancing eBooks and Physical Books Reaction Rates Chemistry Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rates Chemistry Packet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reaction Rates Chemistry Packet Answers
 - Setting Reading Goals Reaction Rates Chemistry Packet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reaction Rates Chemistry Packet Answers
 - Fact-Checking eBook Content of Reaction Rates Chemistry Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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