Long answer type

1. i) The rate law is given by, r = [A] 1/2[B] 2. What is the order of this reaction.

ii) A first order reaction is found to have a rate constant $k = 5.5 \times 10-145 - 1$ Find the Half life of the reaction. iii) Write two differences between order and molecularity.

i)A reaction is 50% complete in 20 minutes at 300K and the same reaction again complete in 5 minutes at 350K. Calculate the activation energy if it is first order reaction [R=8.314 JK-1Mole-1 ; log4 = 0.602]. ii)Define rate constant and order of a reaction

3. i)A first order reaction has a rate constant of 0.0051min-1. If we begin with 0.10M concentration of the reactant, what concentration of the reactant will be left after three hour?

ii)A reaction is second order with respect to a reactant .How is its rates effected when concentration of the reactant is

i) doubled ii)reduced to half

 i)Show graphically how the rate of a 1st order reaction with only one reactant depends upon the concentration of the reactant

ii)Decomposition of phosphine (PH3) proceeds according to the following equation 4PH3(g)→P4(g)+6H2(g). It is found that the reaction follows the following rate equation Rate-k (PH3) The half life of (PH3) is 37.9s at1200C.

i)How much time is required for 3/4th of phosphine to decompose.

ii)What fraction of original sample of PH3reemains behind after I minute. Determine the order of the reaction with respect to A and with respect to B and overall order of the reaction.

5(a) A reaction is second order in A and first order in

B. (i). Write the differential rate equation.

(iii). How is the rate affected on increasing the concentrations of A three times?

(iii). How is the rate affected when the concentrations of both A and B are doubled?

(b). A first order reaction takes 40 minutes for 30% decomposition. Calculate Expt. No [A] mole/I [B] mole/I Rate of formation of D mole/I/min 1 0.1 0.1 5.0x10-3 2 0.3 0.2 6.0x10-2 3 0.3 0.4 2.4x10-1 4 0.4 0.1 2.0x10-2 t1/2 for this reaction. (Given log 1.428 = 0.1548)

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- 1. Understanding the eBook Reaction Rate Review Answers
 - The Rise of Digital Reading Reaction Rate Review Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Reaction Rate Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rate Review Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Reaction Rate Review Answers
 - Personalized Recommendations
 - Reaction Rate Review Answers User Reviews and Ratings
 - Reaction Rate Review Answers and Bestseller Lists
- 5. Accessing Reaction Rate Review Answers Free and Paid eBooks
 - Reaction Rate Review Answers Public Domain eBooks
 - Reaction Rate Review Answers eBook Subscription Services

- Reaction Rate Review Answers Budget-Friendly Options
- 6. Navigating Reaction Rate Review Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Reaction Rate Review Answers Compatibility with Devices
 - Reaction Rate Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Rate Review Answers
 - Highlighting and Note-Taking Reaction Rate Review Answers
 - Interactive Elements Reaction Rate Review Answers
- 8. Staying Engaged with Reaction Rate Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rate Review Answers
- 9. Balancing eBooks and Physical Books Reaction Rate Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rate Review Answers
- 10. Overcoming Reading Challenges
 - \circ Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Rate Review Answers
 - Setting Reading Goals Reaction Rate Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Rate Review Answers
 - Fact-Checking eBook Content of Reaction Rate Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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