

Reactions Practice

Name: _____

Date: _____

Hour: _____

I. Single Replacement Reactions

(Remember, some reactions will NOT proceed. If there is no reaction, write "NR" to the right of the arrow.)

1. Write the reaction for each of the following reactions:

a) Sodium sulfate reacts with magnesium to produce magnesium sulfate and sodium

b) Copper(I) carbonate reacts with aluminum to form copper metal and aluminum carbonate.

2. Complete the following reactions:

a) $\text{Ca}(\text{NO}_3)_2 + \text{Al} \rightarrow$

b) $\text{Zn} + \text{K}_3\text{PO}_4 \rightarrow$

II. Double Replacement Reactions

3. Write the reaction for each of the following reactions:

a) Iron(II) phosphate reacts with sodium chloride to produce iron(II) chloride and sodium phosphate.

b) Calcium carbonate reacts with lithium hydroxide to produce calcium hydroxide and lithium carbonate.

4. Complete the following reactions:

a) $\text{Mg}(\text{NO}_3)_2 + \text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3 \rightarrow$

b) $\text{Ba}(\text{OH})_2 + \text{Na}_3\text{PO}_4 \rightarrow$

III. Combustion Reactions

You can try the following formula to balance: $(\text{C}_x\text{H}_y + (\frac{x+y}{4})\text{O}_2 \rightarrow x\text{CO}_2 + y\text{H}_2\text{O})$

5. Write the combustion reactions for the following compounds.

a) C_2H_4

b) C_8H_{16}

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Judd E. Hollander

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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