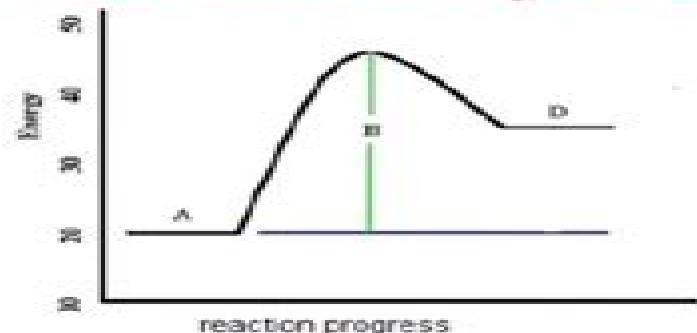


1. Suppose two molecules that can react collide. What **3** circumstances need to occur for the molecules to react?
1. There must be enough energy, 2. They must collide, 3. They must have the correct orientation
2. What does the **activation energy** for a chemical reaction mean?
This is the amount of energy required to make the molecules collide so that a reaction may occur.
3. What role does the **reactivity of the reactants** play in determining the rate of a chemical reaction?
If reactants are not reactive, the reaction may not occur.
4. Using the **collision theory**, explain why sugar sprinkled over a Bunsen burner reacts more quickly than a **chunk** of the same solid.
When the sugar is broken up into individual crystals, there is more surface area and more molecules available to collide and react.
5. Use the **collision theory** to explain why **increasing the concentration** of a reactant usually increases the reaction rate.
Increasing the concentration increases the number of molecules which increases the likelihood that collisions will occur.
6. Use the **collision theory** to explain why **increasing the temperature** usually increases the reaction rate.
Increasing the temperature increases the speed of the molecules. Faster moving molecules collide more.
7. Explain how a **catalyst** affects the activation energy for a chemical reaction.
A catalyst ensure that molecules will collide in the correct orientation, so less energy is needed.
8. **On** the accompanying energy diagram, label the following items:
 - reactants **A**
 - products **D**
 - activation energy **B**
9. **Draw and label** what this diagram would look like if a **catalyst** was added to the reaction.
10. Graph reading
 - a. How much energy do the reactants have?
20 kJ
 - b. How much energy do the products have?
35 kJ
 - c. How much energy is required to activate this un-catalyzed reaction?
Approximately 25 kJ.
 - d. Will this reaction need to absorb energy or will it release energy to the environment? How do you know?
It will absorb energy because the products have more energy than the reactants.



Reaction Rates Worksheet With Answers

G Thomas

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Reaction Rates Worksheet With Answers

1. Understanding the eBook Reaction Rates Worksheet With Answers
 - The Rise of Digital Reading Reaction Rates Worksheet With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates Worksheet With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Reaction Rates Worksheet With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates Worksheet With Answers
 - Personalized Recommendations
 - Reaction Rates Worksheet With Answers User Reviews and Ratings

- Reaction Rates Worksheet With Answers and Bestseller Lists
- 5. Accessing Reaction Rates Worksheet With Answers Free and Paid eBooks
 - Reaction Rates Worksheet With Answers Public Domain eBooks
 - Reaction Rates Worksheet With Answers eBook Subscription Services
 - Reaction Rates Worksheet With Answers Budget-Friendly Options
- 6. Navigating Reaction Rates Worksheet With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Reaction Rates Worksheet With Answers Compatibility with Devices
 - Reaction Rates Worksheet With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Rates Worksheet With Answers
 - Highlighting and Note-Taking Reaction Rates Worksheet With Answers
 - Interactive Elements Reaction Rates Worksheet With Answers
- 8. Staying Engaged with Reaction Rates Worksheet With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rates Worksheet With Answers
- 9. Balancing eBooks and Physical Books Reaction Rates Worksheet With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rates Worksheet With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Rates Worksheet With Answers
 - Setting Reading Goals Reaction Rates Worksheet With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Rates Worksheet With Answers
 - Fact-Checking eBook Content of Reaction Rates Worksheet With Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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