

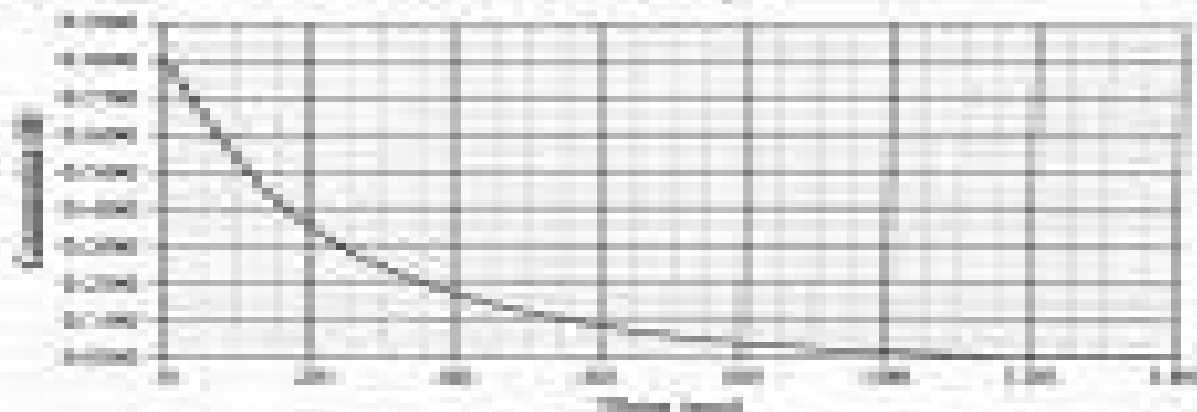
Rate of Reaction

What is the speed of a reaction measured?

Why?

Chemical reactions occur at different speeds. Some are ultrafast (explosions), others very slow. For example, you can leave sodium in your kitchen-drawers with no significant reaction depending on the conditions. In order to study reactions that change the speed of a reaction, we must develop an understanding of how the rate of reaction is measured during a reaction.

Model 1 -- Concentration versus Time Graphs



1. The graph in Model 1. How does the concentration of a specific chemical species change over time?
 - a. What unit is the concentration measured in?
(The concentration is measured in molarity.)
 - b. What unit is the time measured in?
(The time is measured in seconds.)
2. Consider the data in Model 1.
 - a. What specific concentration of the species when the chemical reaction was initiated?
(The initial concentration of the chemical species in the mixture was 0.100 M.)
 - b. Did the concentration of the species increase or decrease over time?
(The concentration of the chemical species decreased over time.)
 - c. Was the species a reactant or product in the reaction? Justify your reasoning.
(The chemical species was a reactant because it was used up during the reaction. Its concentration went down.)

Reaction And Rates Pogil Answer Key

Léo Pomar



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