

Unit 5 Review Sheet – Stoichiometry

Writing Formulas:

1. Cesium Sulfate Cs_2SO_4
2. Copper (II) chlorite $\text{Cu}(\text{ClO}_2)_2$
3. Zinc hydroxide $\text{Zn}(\text{OH})_2$
4. Gold (I) oxide Au_2O
5. Dinitrogen trioxide N_2O_3
6. Phosphorus pentiodide PI_5

Naming:

7. BrF_3 bromine trifluoride
8. N_2O dinitrogen monoxide
9. $\text{Mg}(\text{CN})_2$ magnesium cyanide
10. K_3P potassium phosphide
11. Cu_2SO_3 copper (I) sulfite
12. FeCl_3 iron (III) chloride

Balancing Equations:



Writing and Balancing Equations:



Moles:

17. How many grams are in 3.2 moles of Ag?

$$\frac{3.2 \text{ mol Ag}}{1} \times \frac{107.9 \text{ g Ag}}{1 \text{ mol Ag}} = 345.3 \text{ g Ag}$$

18. How many moles are in 5×10^{25} molecules of H_2O ?

$$\frac{5 \times 10^{25} \text{ molecules H}_2\text{O}}{1} \times \frac{1 \text{ mol H}_2\text{O}}{6.02 \times 10^{23} \text{ molecules H}_2\text{O}} = 83 \text{ mol H}_2\text{O}$$

Stoichiometry:

Use the following balanced equation for #19-21:



19. How many moles of K are needed to react with 3.5 moles of O_2 ?

$$\frac{3.5 \text{ mol O}_2}{1} \times \frac{4 \text{ mol K}}{1 \text{ mol O}_2} = 14 \text{ mol K}$$

20. How many grams of K_2O will be produced when 86.4 g of K react?

$$\frac{86.4 \text{ g K}}{1} \times \frac{1 \text{ mol K}}{39.1 \text{ g K}} \times \frac{2 \text{ mol K}_2\text{O}}{4 \text{ mol K}} \times \frac{94.2 \text{ g K}_2\text{O}}{1 \text{ mol K}_2\text{O}} = 104.1 \text{ g K}_2\text{O}$$

Physics Unit 5 Review Sheet

Leia Levy



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