

Phosphoric acid,  $\text{H}_3\text{PO}_4$ , can be manufactured by a two step process involving the reaction of phosphorus with oxygen, followed by a reaction with water.



- (a) Explain why  $\Delta H_1$  represents the enthalpy change of formation of  $\text{P}_4\text{O}_{10}$ .

.....  
 .....  
 ..... [2]

- (b) Enthalpy changes of formation are shown in the table below.

| Substance                           | Enthalpy change of formation, $\Delta H_f / \text{kJ mol}^{-1}$ |
|-------------------------------------|-----------------------------------------------------------------|
| $\text{P}_4\text{O}_{10}(\text{s})$ | -2984                                                           |
| $\text{H}_2\text{O}(\text{l})$      | -286                                                            |
| $\text{H}_3\text{PO}_4(\text{l})$   | -1267                                                           |

Calculate the enthalpy change of reaction,  $\Delta H_2$ .

$$\Delta H_2 = \dots\dots\dots \text{kJ mol}^{-1} \quad [3]$$

- (c) Write the overall equation for the manufacture of  $\text{H}_3\text{PO}_4$  from  $\text{P}_4$ .

Use this equation to explain why this process has a 100% atom economy.

.....  
 .....  
 .....  
 ..... [2]

[Total: 7]

Turn over

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### **Table of Contents Physics P4 Ocr 2013 Jan Paper**

1. Understanding the eBook Physics P4 Ocr 2013 Jan Paper
  - The Rise of Digital Reading Physics P4 Ocr 2013 Jan Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Physics P4 Ocr 2013 Jan Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physics P4 Ocr 2013 Jan Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physics P4 Ocr 2013 Jan Paper
  - Personalized Recommendations

- Physics P4 Ocr 2013 Jan Paper User Reviews and Ratings
- Physics P4 Ocr 2013 Jan Paper and Bestseller Lists
- 5. Accessing Physics P4 Ocr 2013 Jan Paper Free and Paid eBooks
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  - Physics P4 Ocr 2013 Jan Paper eBook Subscription Services
  - Physics P4 Ocr 2013 Jan Paper Budget-Friendly Options
- 6. Navigating Physics P4 Ocr 2013 Jan Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - Physics P4 Ocr 2013 Jan Paper Compatibility with Devices
  - Physics P4 Ocr 2013 Jan Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physics P4 Ocr 2013 Jan Paper
  - Highlighting and Note-Taking Physics P4 Ocr 2013 Jan Paper
  - Interactive Elements Physics P4 Ocr 2013 Jan Paper
- 8. Staying Engaged with Physics P4 Ocr 2013 Jan Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physics P4 Ocr 2013 Jan Paper
- 9. Balancing eBooks and Physical Books Physics P4 Ocr 2013 Jan Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physics P4 Ocr 2013 Jan Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physics P4 Ocr 2013 Jan Paper
  - Setting Reading Goals Physics P4 Ocr 2013 Jan Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics P4 Ocr 2013 Jan Paper
  - Fact-Checking eBook Content of Physics P4 Ocr 2013 Jan Paper

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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