#### Sample Test for Unit VII

Potentially useful equations:

$$Q=mc\Delta T$$
  $Q=mL$   $Q=mH_c$   $\varepsilon_c=1-\frac{T_C}{T_H}$   $\varepsilon=\frac{W}{Q_m}$ 

Specific heat for water =  $4186 \,\mathrm{J \, kg^{-1}C^{\circ -1}}$ 

1. a) Define the kilocalorie.

b) How much heat energy in kilocalories is required to change 2.3 kg of water from 38°C to 45°C?

2. a) Define Heat of Combustion.

b) A fuel-oil burning boiler is used to heat water from  $20^{\circ}$ C to  $60^{\circ}$ C. The Heat of Combustion for fuel-oil is  $4.19 \times 10^{7}$  J/kg. The boiler is 40% efficient. How much fuel-oil is required to heat 5000 kg of water?

- Fifty grams of glycerine at 35.0°C is mixed with 1.00 kg of water at 20.0°C. Find the specific heat of glycerine if the equilibrium temperature is measured to be 20.42°C; assume no heat losses or gains.
- 4. a) Sketch a typical phase diagram with two phase changes, labeling the axes and transitions. (Also think about what determines the slopes of the T-Q lines: How are they related to the specific heat capacity?)

b) Find the total heat energy required to change 2.0 kg of ice at  $-10^{\circ}$ C into liquid water at 20°C. (Specific heat for ice = 2100J kg<sup>-1</sup>C<sup>-1</sup>, Latent Heat of Fusion =  $3.34 \times 10^{5}$  J/kg, Specific heat for water = 4186 J kg<sup>-1</sup>C<sup>-1</sup>, Latent Heat of Vaporization =  $22.6 \times 10^{5}$  J/kg).

- 5. Define and provide an example for each of the three methods of heat transfer.
- 6. a) State the First Law of Thermodynamics.
  - b) State the Second Law of Thermodynamics.
- 7. a) Draw and label a schematic diagram of a heat engine.

b) Determine the Carnot efficiency of a steam turbine engine which operates between a high temperature of 550°C and a low temperature of 90°C.

#### Short Answers:

a) The keal is the heat energy needed to raise 1 kg of water by 1 C° (from 14.5°C to 15.5°C).

b)  $Q = mc\Delta T$ , where  $c_{\text{water}} = 1 \, \text{kcal} \, \text{kg}^{-1} \text{C}^{\circ -1} (= 41861 \, \text{J} \, \text{kg}^{-1} \text{C}^{\circ -1})$ , so that  $Q = (2.3 \, \text{kg}) (1 \, \text{kcal} \, \text{kg}^{-1} \text{C}^{\circ -1}) (45^{\circ} \text{C} - 38^{\circ} \text{C}) = 16.1 \, \text{kcal}$ .

# **Physics Unit Vii 5 Answers**

Stuart Redman, Lawrence J. Zwier

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