



Physics Thermodynamics

The study of the thermal energy of objects, how it is exchanged between objects, and how it relates to work

Introduction

KEY VARIABLES

Variable	Description	Units
P	Pressure	Pa (N/m ²)
V	Volume	m ³
T	Temperature	K
n	Number of moles in a system	
N	Number of particles in a system	
f	Degrees of freedom	
U	Internal energy of a system	J
Q	Heat added to a system ($Q > 0$)	J
E_{th}	Thermal energy of a system	J
W	Work done by a system ($W > 0$)	J
S	Entropy	J/K
H	Enthalpy	J
F	Helmholtz free energy	J
G	Gibbs free energy	J
ρ	Density	kg/m ³

CONSTANTS

Constant	Symbol & Value
Universal gas constant	$R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1}$
Avogadro's number	$N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$
Boltzmann constant	$k_B = R / N_A = 1.38 \times 10^{-23} \text{ J/K}$
Standard temperature	0°C
Standard pressure	10 ⁵ Pa
Planck's constant	$h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$
Speed of light	$c = 3.00 \times 10^8 \text{ m/s}$
Unit of electric charge	$e = 1.60 \times 10^{-19} \text{ C}$

Conventions

- **Italic subscript:** Indicate which variable is held constant (e.g., C_p means that pressure is held constant to calculate C_p)
- **Roman subscript:** Specify special cases for variables.
- **Subscripts i and f:** Indicate initial and final values for a given variable

Extensive vs. Intensive Properties

- **Extensive properties:** Properties of a system proportional to the extent of the system

EX: Double system size re-volume and mass are doubled. These are extensive variables.

- **Intensive properties:** Properties of a system independent of the extent of the system

EX: Double system size re-pressure and temperature remain the same. These are intensive variables.

Moles

- **A mole of ¹²C atoms** has a mass of 1 g.
 - 1 mole = 6.022×10^{23}
- **Atomic number:** Number of protons in an atom
- **Mass number:** Number of protons and neutrons in an atom
- **Atomic mass:** The weighted average mass of a given atom at rest based on the relative natural abundance of isotopes of that atom
- **Molar mass (M):** Number of grams per mole of an atom or molecule
 - Mass (g) = Mn
- **Molar volume:** Volume occupied by one mole of substance; depends on pressure and temperature

Thermal Energy & Temperature

Definitions

Thermal vs. Internal Energy

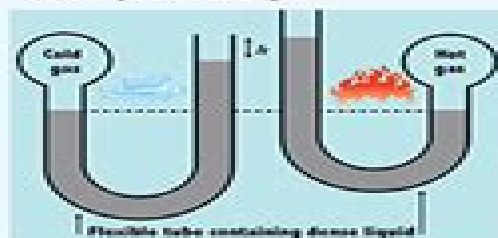
- **Thermal energy (E_{th}):** Energy contained in the random (incoherent) motion of the particles that make up matter
 - Motion can be translational, rotational, and/or vibrational.
 - Nonrandom (coherent) motion does not contribute to thermal energy.
- **Internal energy (U):** Thermal energy plus other forms of energy stored in particles that make up a system (e.g., energy in chemical bonds, nuclei, and mass)

"Heat" vs. "To Heat"

- **Heat:** Thermal energy that is transferred from one system to another
- **To heat:** To increase the temperature of a system

Temperature (T)

- Directly proportional to thermal energy
 - A higher temperature corresponds to higher thermal energy, and vice versa.
 - In layman's terms, higher temperatures correspond to hotter objects and lower temperatures to colder objects.
- **Ideal gas thermometer:** Measures temperature accurately; see image below
 - An ideal gas contains point-like particles that interact only through elastic collisions.
 - The volume of cold gas is heated and the flexible tube is raised so that the hot gas occupies the same volume as the cold gas.
 - The hot gas is at a higher pressure than the cold gas and forces the liquid at height h higher up the tube. This height is proportional to the absolute temperature of the gas.



Thermodynamic State Variables (P , V , T & n)

- Functions of these variables are also thermodynamic state variables.
- The value of state variables does not depend on how the system got into the particular state.

Temperature Scales

"Absolute zero" temperature is the temperature of minimum possible thermal energy.

Arbitrary Scales

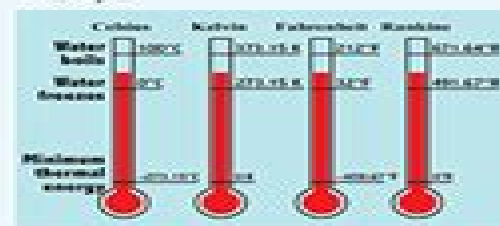
- **Celsius:** Defines absolute zero to be -273.15°C and the triple point of pure water to be 273.16°C .
 - The freezing (boiling) point of water at 1 standard atmosphere of pressure is very close to 0°C (100°C).
- **Fahrenheit:** Defines absolute zero to be -459.67°F and the triple point of pure water to be 32.018°F .
 - The freezing (boiling) point of water at 1 standard atmosphere of pressure is very close to 32°F (212°F).

Absolute Scales

- **Kelvin:** Assigns 0°K as the temperature of minimum possible thermal energy
 - $1^\circ\text{K} = 1^\circ\text{C}$
 - The degree sign is not used for the Kelvin scale.
- **Rankine:** Assigns 0°R as the temperature of minimum possible thermal energy
 - $1^\circ\text{R} = 1^\circ\text{F}$
 - NIST recommends against using the Rankine scale.

Conversion between Temperature Scales

- $T_C = \frac{5}{9}(T_F - 32^\circ)$
- $T_F = \frac{9}{5}T_C + 32^\circ$
- $T_K = T_C + 273.15$
- $T_R = T_F + 459.67$
- $T_K = \frac{5}{9}T_R$



Laws of Thermodynamics

Zeroth Law of Thermodynamics

- Two systems in thermal contact at the same temperature are in thermal equilibrium.
- Two systems in thermal equilibrium with a third system are in thermal equilibrium with each other.

First Law of Thermodynamics

- The energy of a system is conserved.
- $\Delta U = Q - W$
 - The change in the internal energy (ΔU) of a system is the difference between the amount of heat (Q) supplied to the system and the amount of work (W) done by the system.

Second Law of Thermodynamics

Entropy

- **Entropy:** The quantitative measure of the ability of thermal energy to do work per unit temperature
 - Entropy (S) is a state variable.
 - A high-entropy system can do more work than a low-entropy system at the same temperature.
 - Entropy cannot be measured, only calculated.
 - Entropy is an extensive property, so the entropy of two systems A and B is the sum of entropies: $S_{total} = S_A + S_B$
- **Macrostate:** State of a system described by state variables T , P , V , and n

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Raffaela Di Napoli



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Table of Contents Physics Thermodynamics Study Guide

1. Understanding the eBook Physics Thermodynamics Study Guide
 - The Rise of Digital Reading Physics Thermodynamics Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics Thermodynamics Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physics Thermodynamics Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Thermodynamics Study Guide

- Personalized Recommendations
- Physics Thermodynamics Study Guide User Reviews and Ratings
- Physics Thermodynamics Study Guide and Bestseller Lists
- 5. Accessing Physics Thermodynamics Study Guide Free and Paid eBooks
 - Physics Thermodynamics Study Guide Public Domain eBooks
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 - ePub, PDF, MOBI, and More
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 - Highlighting and Note-Taking Physics Thermodynamics Study Guide
 - Interactive Elements Physics Thermodynamics Study Guide
- 8. Staying Engaged with Physics Thermodynamics Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics Thermodynamics Study Guide
- 9. Balancing eBooks and Physical Books Physics Thermodynamics Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Thermodynamics Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physics Thermodynamics Study Guide
 - Setting Reading Goals Physics Thermodynamics Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics Thermodynamics Study Guide

- Fact-Checking eBook Content of Physics Thermodynamics Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web apr 20 2020 known alternatively as the lowbrow movement pop surrealism is an art form that originated in la s underground scene in the 1970s like other surrealist art forms lowbrow art strives to reach deep into the unconscious mind and bring to

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web jun 7 2023 pop surrealism or lowbrow art is a movement that arose in los angeles in the late 1970s the movement is young and what could be considered the founding generation are only in their fifties or early sixties

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web jul 1 2021 what sets pop surrealism also known as lowbrow art apart is the incorporation of elements from popular and mass culture sometimes ironic sometimes playful lowbrow art features everything from notable brands and memorable characters to cultural trends and hot takes on current events

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against cultural and political rationalism surrealism grew out of the dada movement but its emphasis was on positive
expression members included salvador dali rene magritte meret oppenheim and leonora carrington

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lowbrow into the fine arts 2 marion peck marion peck received a bfa from rhode island school of design in 1985 at age 18
subsequently she 3 amy sol amy sol is an

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were considered outlaws by critics at the time this style of art bridges the gap between surrealism and pop culture taking the
best from both worlds to create satirical works submerged in fantasy

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techniques in pop surrealism 1 realism with a twist 2 metamorphosis and transformation 3 symbolism and allegory influential

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