

THERMO DYNAMICS

1. **System:**
A particular portion of matter is called a system
2. **Surroundings:**
Every thing outside of the system that may have a direct influence on the system is referred as the surroundings.
3. **Types of systems:**
There are 3 types of systems
 - a. **Open system:**
A system which exchanges both energy and matter with its surroundings is called as open system
 - b. **Closed system:**
A system which can exchange energy only and not matter with surrounding is known as closed system.
 - c. **Isolated system:**
A system which can exchange neither energy nor matter with the surroundings is called isolated system.
Examples: hot box, thermo flask
4. **Thermo dynamics co-ordinates (or) variables:**
The parameters like pressure, volume, temperature entropy, internal energy, work, etc characterising. The state of the system are called thermo dynamic variables
5. **Thermo dynamic system:**
A system that can be described in terms of the thermo dynamic co-ordinates is called thermo dynamic system:
6. **Heat:**
Heat is a form of energy which transfer from a body at higher temperature to a body at lower temperature.
Units: Unit of heat in SI - Joule
Unit of heat in C. G. S - Calorie
 $1 \text{ Cal} = 4.2 \text{ J}$
7. **Thermal equilibrium:**
Two bodies are said to be in thermal equilibrium if there is no exchange of heat energy between them when connected thermally with each other.
8. **Zeroth law of thermo dynamics:**
If two systems A and B are in thermal equilibrium with third system C separately then they (A, B) must be in thermal equilibrium with each other. Mathematical representation

$$f_A(P_A, V_A) = f_B(P_B, V_B) = f_C(P_C, V_C) = T$$
 (or)

$$f(P, V, T) = 0$$
9. **Calorie:**
The quantity of heat required to rise the tem-

perature of 1gm of water through 1°C is called 1 calorie.

10. **Standard Calorie:**

The quantity of heat required to rise the temperature of 1gm of water from 14.5°C to 15.5°C is called standard calorie

11. **External work:**

When the force is exerted by the system on its surroundings produces a displacement the work done by the system is called external work.

(or)

When the force is exerted by the surroundings on the system produces a displacement the work done by the surroundings is called external work.

- a. The work done by the system on the surroundings is taken as '+ve'
- b. The work done by the surroundings on the system is taken as '-ve'

12. **Internal work:**

When the work done by one part of the system on another part of the same system is called internal work.

13. **Internal energy:**

Internal energy of a thermo dynamic system is defined as the energy possessed by the system by virtue of its molecular constitution and the position of the molecules

- a. In a gas the molecules are in the state of random motion hence they possess kinetic energy.
- b. Due to molecular forces of attraction among the molecules they possess potential energy.
- c. The sum of potential energy and kinetic energy of individual molecules in the system is called internal energy.
Internal energy $U = P.E. + K.E$
 $\therefore$ change in internal energy

$$dU = U_f - U_i$$
 where U_f is final internal energy
 U_i is initial internal energy
- d. Internal energy of a system increases with temperature.

14. **Heat capacity (or) thermal capacity**

The quantity of heat required to rise the temperature of a substance through 1°C is called heat capacity it is represented by 'H'

$$\therefore H = \frac{\text{Change in heat energy}}{\text{rise in temperature}}$$

$$H = \frac{dQ}{dT}$$

Physics Thermodynamics Chapter Test

Guillaume Favre



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Table of Contents Physics Thermodynamics Chapter Test

1. Understanding the eBook Physics Thermodynamics Chapter Test
 - The Rise of Digital Reading Physics Thermodynamics Chapter Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics Thermodynamics Chapter Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physics Thermodynamics Chapter Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Thermodynamics Chapter Test
 - Personalized Recommendations
 - Physics Thermodynamics Chapter Test User Reviews and Ratings
 - Physics Thermodynamics Chapter Test and Bestseller Lists
5. Accessing Physics Thermodynamics Chapter Test Free and Paid eBooks
 - Physics Thermodynamics Chapter Test Public Domain eBooks
 - Physics Thermodynamics Chapter Test eBook Subscription Services
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6. Navigating Physics Thermodynamics Chapter Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Physics Thermodynamics Chapter Test Compatibility with Devices
 - Physics Thermodynamics Chapter Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics Thermodynamics Chapter Test
 - Highlighting and Note-Taking Physics Thermodynamics Chapter Test
 - Interactive Elements Physics Thermodynamics Chapter Test
8. Staying Engaged with Physics Thermodynamics Chapter Test

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Physics Thermodynamics Chapter Test
- 9. Balancing eBooks and Physical Books Physics Thermodynamics Chapter Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Thermodynamics Chapter Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physics Thermodynamics Chapter Test
 - Setting Reading Goals Physics Thermodynamics Chapter Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics Thermodynamics Chapter Test
 - Fact-Checking eBook Content of Physics Thermodynamics Chapter Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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