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Chapter 1

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Answers to "Why?" in the text

Page 31:

Oxygen has an atomic mass of 16 whereas it is 14 for nitrogen. The O_2 molecule is therefore heavier than the N_2 molecule. Thus, from $\frac{1}{2}mv^2 = 3(\frac{1}{2}kT)$, the rms velocity of O_2 molecules is smaller than that of N_2 molecules.

Page 34, footnote 11

For small extensions, the difference between the engineering and instantaneous strains due to a temperature change are the same. Historically, mechanical and civil engineers measured extension by monitoring the change in length, ΔL ; and the instantaneous length L was not measured. It is not trivial to measure both the instantaneous length and the extension simultaneously. However, since we know L_o and measure ΔL , the instantaneous length $L = L_o + \Delta L$. Is the difference important? Consider a sample of length L_o that extends to a final length L due to a temperature change from T_o to T . Let $\epsilon = (L - L_o) / L_o = \Delta L / L_o$ be the engineering strain.

The engineering definition of strain and hence the thermal expansion coefficient is

$$\text{Engineering strain} = \frac{\delta L}{L_o} = \lambda \delta T$$

so that thermal expansion from T_o to T gives,

$$\int_{L_o}^L \frac{dL}{L_o} = \int_{T_o}^T \lambda dT \quad \therefore \quad \frac{\Delta L}{L_o} = \lambda(T - T_o) \quad \therefore \quad \epsilon = \lambda(T - T_o) \quad (1)$$

where $\epsilon = \Delta L / L_o$ is the engineering strain as defined above

Physics definition of strain and hence the thermal expansion coefficient is

$$\text{Instantaneous strain} = \frac{\delta L}{L} = \lambda \delta T$$

so that thermal expansion from T_o to T gives,

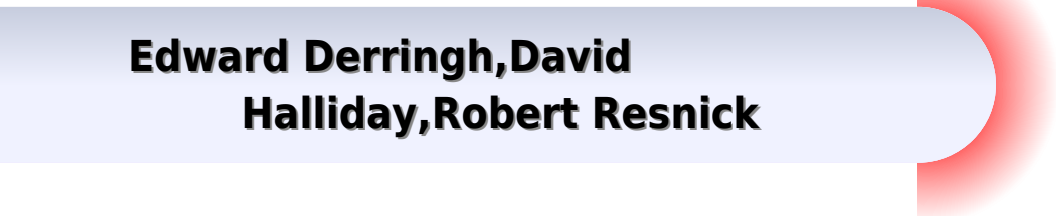
$$\int_{L_o}^L \frac{dL}{L} = \int_{T_o}^T \lambda dT \quad \therefore \quad \ln\left(\frac{L}{L_o}\right) = \lambda(T - T_o) \quad \therefore \quad \ln(1 + \epsilon) = \lambda(T - T_o) \quad (2)$$

We can expand the $\ln(1 + \epsilon)$ term for small ϵ , so that Equation (2) essentially becomes Equation (1)

1.1 Virial theorem The Li atom has a nucleus with a $+3e$ positive charge, which is surrounded by a full $1s$ shell with two electrons, and a single valence electron in the outer $2s$ subshell. The atomic radius of the

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**Edward Derrin, David
Halliday, Robert Resnick**



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Table of Contents Physics Solutions Manual Chapter

1. Understanding the eBook Physics Solutions Manual Chapter
 - The Rise of Digital Reading Physics Solutions Manual Chapter
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics Solutions Manual Chapter
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physics Solutions Manual Chapter
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Solutions Manual Chapter
 - Personalized Recommendations

- Physics Solutions Manual Chapter User Reviews and Ratings
- Physics Solutions Manual Chapter and Bestseller Lists
- 5. Accessing Physics Solutions Manual Chapter Free and Paid eBooks
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 - Physics Solutions Manual Chapter Budget-Friendly Options
- 6. Navigating Physics Solutions Manual Chapter eBook Formats
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 - Physics Solutions Manual Chapter Compatibility with Devices
 - Physics Solutions Manual Chapter Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics Solutions Manual Chapter
 - Highlighting and Note-Taking Physics Solutions Manual Chapter
 - Interactive Elements Physics Solutions Manual Chapter
- 8. Staying Engaged with Physics Solutions Manual Chapter
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics Solutions Manual Chapter
- 9. Balancing eBooks and Physical Books Physics Solutions Manual Chapter
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Solutions Manual Chapter
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physics Solutions Manual Chapter
 - Setting Reading Goals Physics Solutions Manual Chapter
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics Solutions Manual Chapter
 - Fact-Checking eBook Content of Physics Solutions Manual Chapter

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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