

Physics

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(Chapter 13) (Kinetic Theory) (Additional Exercises)
(Class 11)

Question 13.14:

Given below are densities of some solids and liquids. Give rough estimates of the size of their atoms:

Substance	Atomic Mass (u)	Density (10^3 Kg m^{-3})
Carbon (diamond)	12.01	2.22
Gold	197.00	19.32
Nitrogen (liquid)	14.01	1.00
Lithium	6.94	0.53
Fluorine (liquid)	19.00	1.14

[Hint: Assume the atoms to be 'tightly packed' in a solid or liquid phase, and use the known value of Avogadro's number. You should, however, not take the actual numbers you obtain for various atomic sizes too literally. Because of the crudeness of the tight packing approximation, the results only indicate that atomic sizes are in the range of a few Å].

Answer 13.14:

Substance	Radius (Å)
Carbon (diamond)	1.29
Gold	1.59
Nitrogen (liquid)	1.77
Lithium	1.73
Fluorine (liquid)	1.88

Atomic mass of a substance = M and Density of the substance = ρ

Avogadro's number = $N = 6.023 \times 10^{23}$

Volume of N number of molecules = $\frac{4}{3} \pi r^3 N$... (i)

Volume of one mole of a substance = $\frac{M}{\rho}$... (ii)

$$\frac{4}{3} \pi r^3 N = \frac{M}{\rho} \Rightarrow r = \sqrt[3]{\frac{3M}{4\pi\rho N}}$$

For gold: $M = 197.00 \times 10^{-3} \text{ kg}$ $\rho = 19.32 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 197 \times 10^{-3}}{4\pi \times 19.32 \times 10^3 \times 6.23 \times 10^{23}} \right)^{\frac{1}{3}} = 1.77 \text{ Å}$$

Hence, the radius of a liquid nitrogen atom is 1.77 Å.

For lithium: $M = 6.94 \times 10^{-3} \text{ kg}$ $\rho = 0.53 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 6.94 \times 10^{-3}}{4\pi \times 0.53 \times 10^3 \times 6.23 \times 10^{23}} \right)^{\frac{1}{3}} = 1.73 \text{ Å}$$

Hence, the radius of a lithium atom is 1.73 Å.

For liquid fluorine: $M = 19.00 \times 10^{-3} \text{ kg}$ and $\rho = 1.14 \times 10^3 \text{ kg m}^{-3}$

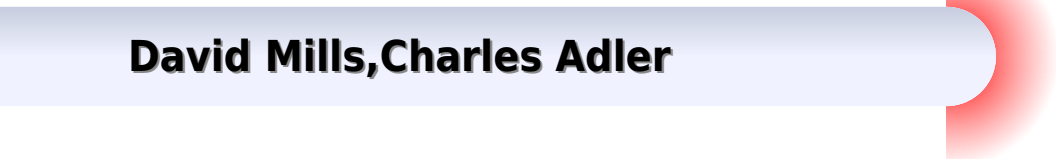
$$\therefore r = \left(\frac{3 \times 19 \times 10^{-3}}{4\pi \times 1.14 \times 10^3 \times 6.023 \times 10^{23}} \right)^{\frac{1}{3}} = 1.88 \text{ Å}$$

Hence, the radius of a liquid fluorine atom is 1.88 Å.

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Physics Principles Problems Answers Chapter 13

David Mills, Charles Adler



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