

A container has 0.50 kg of ice at 0.00 °C. A tube connected from a boiler let 0.100 kg of steam at 100.0 °C inside the container. What is the temperature of the system when it has reach equilibrium? Assume that no heat is absorbed by the container.

Given: $m_{ice} = 0.50 \text{ kg}$

$m_{steam} = 0.100 \text{ kg}$ steam

$L_f \text{ water} = 33.5 \times 10^4 \text{ J/kg}$

$t_i = 0.00^\circ\text{C}$

$t_i = 100^\circ\text{C}$

$L_v \text{ water} = 22.6 \times 10^5 \text{ J/kg}$

$t_f = ?$

$t_f = ?$

$C_{water} = 4180 \text{ J/kg}\cdot^\circ\text{C}$

① Heat needed to make ice melt

$Q_1 = mL_f$

$Q_1 = (0.50 \text{ kg}) (33.5 \times 10^4 \text{ J/kg})$

$Q_1 = 167500 \text{ J/kg}\cdot^\circ\text{C}$

② Heat needed to make melted ice to warm from 0 to t_f

$Q_2 = mc\Delta T$

$Q_2 = (0.50 \text{ kg}) (4180 \text{ J/kg}\cdot^\circ\text{C}) (t_f - 0.00^\circ\text{C})$

$Q_2 = (2095 \text{ J/}^\circ\text{C}) (t_f)$

③ Total heat needed for ice

$Q_{ice} = Q_1 + Q_2$

④ Heat to needed to form steam to water

$Q_3 = mL_v$

Eq. 1 $Q_{ice} = 167500 \text{ J} + (2095 \text{ J/}^\circ\text{C}) (t_f)$

$Q_3 = (0.100 \text{ kg}) (22.6 \times 10^5 \text{ J/kg}) = 226000 \text{ J}$

⑤ Heat needed to form vaporized steam

from 100 °C to t_f

$Q_4 = mc\Delta T$

$Q_4 = (0.100 \text{ kg}) (4180 \text{ J/kg}\cdot^\circ\text{C}) - (100 - 100^\circ\text{C}) = 0 \text{ J}$

$Q_5 = 4180 \text{ J/}^\circ\text{C} (100^\circ\text{C} - t_f)$

Eq. 2 $Q_5 = 418000 \text{ J} - (4180 \text{ J/}^\circ\text{C}) (t_f)$

⑥ Total heat needed for steam

⑦ The system equilibrium

$Q_{steam} = Q_3 + Q_4$

$Q_{ice} = Q_{steam}$

$Q_{steam} = 226000 \text{ J} + 4180 \text{ J} - (4180 \text{ J/}^\circ\text{C}) (t_f)$

$Q_{ice} = Q_{steam}$

$Q_{ice} = 209500 \text{ J} + (4180 \text{ J/}^\circ\text{C}) (t_f)$

⑧ Equating Eq. 1 and Eq. 2

$167500 \text{ J} + (2095 \text{ J/}^\circ\text{C}) (t_f) = 226000 \text{ J} + (4180 \text{ J/}^\circ\text{C}) (t_f)$

$167500 \text{ J} - 226000 \text{ J} = (4180 \text{ J/}^\circ\text{C}) (t_f) - (2095 \text{ J/}^\circ\text{C}) (t_f)$

$-58500 \text{ J} = (2085 \text{ J/}^\circ\text{C}) (t_f)$

$t_f = \frac{-58500 \text{ J}}{2085 \text{ J/}^\circ\text{C}}$

$t_f = -28^\circ\text{C}$

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