

1. Étude de la molécule de l'acide ascorbique

La molécule d'acide ascorbique est représentée ci-dessous :

1.1. Les propriétés acido-basiques de cette molécule sont dues à l'hydrogène porté par l'oxygène du groupe caractéristique associé à l'atome de carbone en position 3.

Représenter l'ion ascorbate, base conjuguée de l'acide ascorbique.

Il suffit d'enlever l'ion H^+ du groupe OH porté par le carbone n°3

L'ion ascorbate est donc :



qu'on notera pour plus de facilité A^-

2. Vérification de la masse d'acide ascorbique dans un comprimé

On souhaite vérifier l'indication portée sur la boîte concernant la masse d'acide ascorbique présente dans un comprimé, à l'aide d'un titrage acidobasique suivi par conductimétrie.

Une solution aqueuse S_A de volume $V = 200,0 \text{ mL}$, est préparée à partir d'un comprimé entier.

On prélève un volume $V_a = (20,0 \pm 0,1) \text{ mL}$ de la solution aqueuse S_A que l'on titre par une solution aqueuse d'hydroxyde de sodium ($Na^+_{aq} + HO^-_{aq}$).

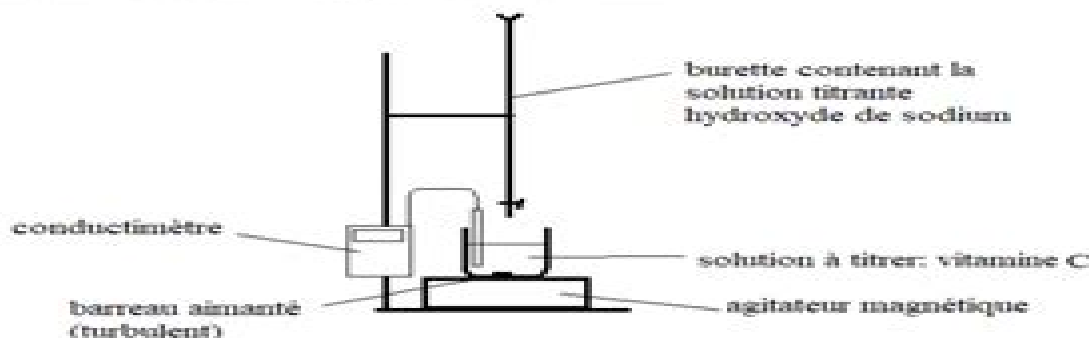
2.1. Rédiger le protocole de préparation de la solution aqueuse S_A .

Le correcteur attend juste que vous disiez que la solution doit être préparée avec une fiole jaugée, les volumes prélevés avec une pipette. On fait un dosage, il faut donc être précis.

2.2. Faire un schéma légendé du dosage conductimétrique.

Dans ce genre de question, aller à l'essentiel, ne faites pas un schéma 3D avec ombres...

Soyez appliqué, mais on ne vous demande pas une œuvre d'art.



Physique Chimie Bac Pondichery 2015

Hélène Hervé, Sonia Madani



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