

# Periodic and Groups: Valence Electrons: Mastery

## Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group					Group	
	1A	2A	3A	4A	5A	6A	7A	8A	18A
1	H 1H	He 2He	Transition metals (not shown here)					Li 3Li	Ne 10Ne
2	B 5B	C 6C						Be 4Be	F 9F
3	Al 13Al	Si 14Si						Na 11Na	Ar 18Ar
4	K 19K	Ca 20Ca						Mg 12Mg	Kr 36Kr

Thinking like the game "Monopoly", answer that an element is in period 4 group

What element is in Group 1A and Period 1?

What group number is in Period 4?

What element is in Group 1, and Period 18A? (Group 1 is not in Group 18A)

What element is in Group 2A and Period 1?

What group number is in Period 2A?

## Valence Electrons

Valence Electrons are the outermost electrons in an atom. Each group contains has the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Lithium 3



3 valence electrons

Hydrogen and Helium are both in Group 1A (Hydrogen 1 and Helium 2)

Group 1A, Lithium 3A and 4A have 4 valence electrons. How many valence electrons will they have?

How many electrons?

**Group 18A** - Elements in group 18A have a full shell of electrons. Elements in group 18A and 18A are called noble gases. Helium, Neon, Argon, Krypton, Xenon, and Radon are noble gases. "He 2, Ne 10, Ar 18, Kr 36, Xe 54, Rn 86" Group 18A has a full shell of valence electrons (noble gases). 18 valence electrons will have, plus, or minus 1 valence electron.

# Periods And Groups Valence Electrons Masses Answers

**B Lingard**



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## **Table of Contents Periods And Groups Valence Electrons Masses Answers**

1. Understanding the eBook Periods And Groups Valence Electrons Masses Answers
  - The Rise of Digital Reading Periods And Groups Valence Electrons Masses Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Periods And Groups Valence Electrons Masses Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periods And Groups Valence Electrons Masses Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periods And Groups Valence Electrons Masses Answers
  - Personalized Recommendations
  - Periods And Groups Valence Electrons Masses Answers User Reviews and Ratings
  - Periods And Groups Valence Electrons Masses Answers and Bestseller Lists
5. Accessing Periods And Groups Valence Electrons Masses Answers Free and Paid eBooks
  - Periods And Groups Valence Electrons Masses Answers Public Domain eBooks
  - Periods And Groups Valence Electrons Masses Answers eBook Subscription Services
  - Periods And Groups Valence Electrons Masses Answers Budget-Friendly Options
6. Navigating Periods And Groups Valence Electrons Masses Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Periods And Groups Valence Electrons Masses Answers Compatibility with Devices
  - Periods And Groups Valence Electrons Masses Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Periods And Groups Valence Electrons Masses Answers
    - Highlighting and Note-Taking Periods And Groups Valence Electrons Masses Answers
    - Interactive Elements Periods And Groups Valence Electrons Masses Answers
  8. Staying Engaged with Periods And Groups Valence Electrons Masses Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Periods And Groups Valence Electrons Masses Answers
  9. Balancing eBooks and Physical Books Periods And Groups Valence Electrons Masses Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Periods And Groups Valence Electrons Masses Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Periods And Groups Valence Electrons Masses Answers
    - Setting Reading Goals Periods And Groups Valence Electrons Masses Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Periods And Groups Valence Electrons Masses Answers
    - Fact-Checking eBook Content of Periods And Groups Valence Electrons Masses Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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