

Periodic Trends Guided Note Sheet

Name _____

Period _____

Date _____

Fill in the blanks. Some blanks have word choices written after them to help you. Refer to the drawings and questions from the first page to help you with the fill in the other pages.

I. Atomic Radius in a Group

- A. As you move down a group, the nucleus becomes _____ (stronger, weaker) because more _____ (electrons, protons) are being added to it. This stronger nucleus can pull the electrons in more tightly, making the atoms **smaller**; **smaller**.
- B. _____ (Protons, Electrons) are also added to the atom but are added to _____ (larger, smaller, the same) energy levels resulting in **larger** atoms.
- C. The increased number of inner electrons shield (or block) the _____ (positive, negative) pull of the nucleus.
- D. So, electron shielding _____ (increases, decreases) causing the electrons on the outermost energy levels to be even farther from the nucleus than expected.
- E. This causes the atoms at the bottom of a Group to be even _____ (larger, smaller) than expected.
- F. Therefore, atomic radius _____ (increases, decreases) as you move down a group.

II. Atomic Radius in a Period

- A. As you move right across a Period, the nucleus becomes _____ (stronger, weaker) because more _____ (electrons, protons) are being added to it.
- B. _____ (Protons, Electrons) are also added but they are added to _____ (the same, a larger, a smaller) energy level.
- C. This causes the amount of electron shielding from electrons on inner energy levels to remain _____ (constant, varied).
- D. Therefore, electrons are pulled in _____ (more, less) tightly because they feel the pull from nucleus more strongly.
- E. This causes the atoms on the right of the Periodic Table to be _____ (larger, smaller) than those on the left.
- F. So, atomic radius _____ (increases, decreases) as you move across a Period from left to right.

III. Sizes of an Ion Versus Sizes of an Atom (Digital Textbook—Section 11.2)

- A. Metal cations are _____ (smaller, larger) than the neutral atoms.
- B. because _____
- C. Nonmetal anions are _____ (smaller, larger) than the neutral atoms.
- D. because _____

IV. Ionization Energy in a Group

- A. Ionization energy is the energy required to _____ the outermost electron from an atom.
- B. As you move down a Group, the nucleus becomes _____ (stronger, weaker) because there are _____ (more, fewer) protons in it.
- C. **Constant**, as you move down a group, there are _____ (fewer, more) energy levels making the atom _____ (larger, smaller).
- D. In atoms at the top of a group, the outermost electrons are closer to the nucleus so the electrons can feel the pull of nucleus _____ (more, less) strongly.
- E. The outermost electrons also feel the pull of the nucleus less strongly because there is _____ (less, more) electron shielding from the _____ (decreased, increased) number of inner electrons.
- F. Therefore, _____ (more, less) energy is required to remove the outermost electron from a larger atom.

Periodic Trends Review Notes Sheet Answers

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Table of Contents Periodic Trends Review Notes Sheet Answers

1. Understanding the eBook Periodic Trends Review Notes Sheet Answers
 - The Rise of Digital Reading Periodic Trends Review Notes Sheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Review Notes Sheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Review Notes Sheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Review Notes Sheet Answers
 - Personalized Recommendations
 - Periodic Trends Review Notes Sheet Answers User Reviews and Ratings
 - Periodic Trends Review Notes Sheet Answers and Bestseller Lists
5. Accessing Periodic Trends Review Notes Sheet Answers Free and Paid eBooks

- Periodic Trends Review Notes Sheet Answers Public Domain eBooks
- Periodic Trends Review Notes Sheet Answers eBook Subscription Services
- Periodic Trends Review Notes Sheet Answers Budget-Friendly Options
- 6. Navigating Periodic Trends Review Notes Sheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Review Notes Sheet Answers Compatibility with Devices
 - Periodic Trends Review Notes Sheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Review Notes Sheet Answers
 - Highlighting and Note-Taking Periodic Trends Review Notes Sheet Answers
 - Interactive Elements Periodic Trends Review Notes Sheet Answers
- 8. Staying Engaged with Periodic Trends Review Notes Sheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Review Notes Sheet Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Review Notes Sheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Review Notes Sheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Review Notes Sheet Answers
 - Setting Reading Goals Periodic Trends Review Notes Sheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Review Notes Sheet Answers
 - Fact-Checking eBook Content of Periodic Trends Review Notes Sheet Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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