# Graphing Periodic Trends Lab

Name\_\_\_\_\_Hr\_\_\_

PURPOSE: To determine how certain properties are periodic when the elements are arranged in periods or groups.

DISCUSSION: There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

#### REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

ionization energy - The energy required to remove an electron from an atom.

#### DATA

Properties of the First 18 Elements

| Properties of the First 18 |                  | Elements                     |                   |
|----------------------------|------------------|------------------------------|-------------------|
| Element                    | Atomic<br>Number | Atomic Radius<br>(Angstroms) | Ionization Energy |
| н                          | 1                | 0.79                         | 314               |
| He                         | 2                | 0.49                         | 567               |
| Li                         | 3                | 2.05                         | 124               |
| Ве                         | 4                | 1.4                          | 215               |
| 8                          | 5                | 1.17                         | 191               |
| C                          | 6                | 0.91                         | 260               |
| M                          | 7                | 0.75                         | 335               |
| 0                          | 8                | 0.65                         | 314               |
| F                          | 9                | 0.57                         | 402               |
| No                         | 10               | 7                            | 2                 |
| Ma                         | 1.1              | 2.23                         | 119               |
| Mg                         | 12               | 1.72                         | 176               |
| A.I                        | 13               | 1.82                         | 138               |
| Si                         | 1.4              | 1.46                         | 188               |
| P                          | 15               | 1.23                         | 242               |
| S                          | 16               | 1.09                         | 239               |
| CI                         | 17               | 0.97                         | 299               |
| Ar                         | 18               | ?                            | 2                 |
|                            |                  |                              |                   |

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with guestions marks.

- Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)
- 2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)

## **Periodic Trends Lab Answers**

Shasha Hu

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